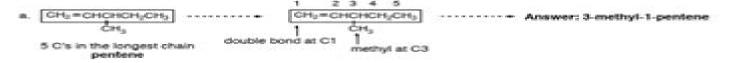
Chapter 13 Unsaturated Hydrocarbons

Solutions to In-Chapter Problems

- 13.1 To draw a complete structure for each condensed structure, first draw in the multiple bonds. Then draw in all the other C's and H's, as in Example 13.1.
- 13.2 To determine whether each molecular formula corresponds to a saturated hydrocarbon, alkene, or alkyne, recall that the formula for a saturated hydrocarbon is C₀H_{2n-2}, the formula for an alkene is C₀H_{2n-2}.
 - a. $C_3H_6=C_nH_{2n}=$ alkene c. $C_6H_{14}=C_nH_{2n-2}=$ alkene d. $C_6H_{12}=C_nH_{2n-2}=$ alkene
- Use the general formulas [saturated hydrocarbon (C_nH_{2n-2}), alkene (C_nH_{2n}), and alkyne (C_nH_{2n-2})] to determine the molecular formula for each compound.
 - a. alkene = C_aH_{2a} , $4 \times 2 = 8$, C_4H_1 b. saturated hydrocarbon = C_aH_{2a-2} , $(6 \times 2) + 2 = 14$, C_6H_{14} c. alkyne = C_aH_{2a-2} , $(7 \times 2) - 2 = 12$, C_7H_{12} d. alkene = C_aH_{2a} , $5 \times 2 = 10$, C_7H_{10}
- 13.4 Give the IUPAC name for each compound using the following steps, as in Example 13.2:
 - [1] Find the longest chain containing both carbon atoms of the multiple bond.
 - [2] Number the chain to give the double bond the lower number.
 - [3] Name and number the substituents and write the complete name.



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Ying-Ying Zheng

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Table of Contents Langfield Smith 6e Chapter 13 Solution

- 1. Understanding the eBook Langfield Smith 6e Chapter 13 Solution
 - The Rise of Digital Reading Langfield Smith 6e Chapter 13 Solution
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Langfield Smith 6e Chapter 13 Solution
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Langfield Smith 6e Chapter 13 Solution
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Langfield Smith 6e Chapter 13 Solution
 - Personalized Recommendations
 - Langfield Smith 6e Chapter 13 Solution User Reviews and Ratings
 - Langfield Smith 6e Chapter 13 Solution and Bestseller Lists

- 5. Accessing Langfield Smith 6e Chapter 13 Solution Free and Paid eBooks
 - Langfield Smith 6e Chapter 13 Solution Public Domain eBooks
 - Langfield Smith 6e Chapter 13 Solution eBook Subscription Services
 - Langfield Smith 6e Chapter 13 Solution Budget-Friendly Options
- 6. Navigating Langfield Smith 6e Chapter 13 Solution eBook Formats
 - o ePub, PDF, MOBI, and More
 - Langfield Smith 6e Chapter 13 Solution Compatibility with Devices
 - Langfield Smith 6e Chapter 13 Solution Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Langfield Smith 6e Chapter 13 Solution
 - Highlighting and Note-Taking Langfield Smith 6e Chapter 13 Solution
 - Interactive Elements Langfield Smith 6e Chapter 13 Solution
- 8. Staying Engaged with Langfield Smith 6e Chapter 13 Solution
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Langfield Smith 6e Chapter 13 Solution
- 9. Balancing eBooks and Physical Books Langfield Smith 6e Chapter 13 Solution
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Langfield Smith 6e Chapter 13 Solution
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Langfield Smith 6e Chapter 13 Solution
 - Setting Reading Goals Langfield Smith 6e Chapter 13 Solution
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Langfield Smith 6e Chapter 13 Solution
 - Fact-Checking eBook Content of Langfield Smith 6e Chapter 13 Solution
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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