

5 (d) Explain why the current in the external circuit of this cell falls to zero after the cell has operated for some time.

.....

.....

(1 mark)

5 (e) The simplified electrode reactions in a rechargeable lithium cell are

Electrode **A** $\text{Li}^+ + \text{MnO}_2 + \text{e}^- \longrightarrow \text{LiMnO}_2$ $E = -0.15 \text{ V}$

Electrode **B** $\text{Li}^+ + \text{e}^- \longrightarrow \text{Li}$

Electrode **B** is the negative electrode.

5 (e) (i) The e.m.f. of this cell is 2.90 V.

Use this information to calculate a value for the electrode potential of electrode **B**.

.....

.....

(1 mark)

5 (e) (ii) Write an equation for the overall reaction that occurs when this lithium cell is being recharged.

.....

.....

.....

(2 marks)

5 (e) (iii) Suggest why the recharging of a lithium cell may lead to release of carbon dioxide into the atmosphere.

.....

.....

(1 mark)

June 2013 Papers As Level Chemistry

DP Hallahan



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