

Intro to Acids and Bases

Name: _____
 Class: _____
 Date: _____

Define each term: Definitions of acids and bases:

Arrhenius Definitions:

- acid substances that release dissolved hydrogen ions (H^+) (acids H^+ gases formed in gases as the hydrogens form H_2 and H_2O (acid H_2O) and conjugated anions)
- base substances which dissociate water molecules (H_2O)

Brønsted-Lowry Definitions:

- acid substances that donate a proton (H^+) (acid donate)
- base substances which accept a proton (H^+) (base receive)

Table 1: Representative reactions for H^+ (acid) and H_2O (base) equilibrium reactions:

Reaction	K_c
1. $HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$	1.0×10^6
2. $HNO_3(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + NO_3^-(aq)$	2.0×10^5
3. $H_2SO_4(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + HSO_4^-(aq)$	1.0×10^3
4. $H_2SO_4(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + HSO_4^-(aq)$	1.0×10^3
5. $H_2O(l) + H_2O(l) \rightleftharpoons H_3O^+(aq) + OH^-(aq)$	1.0×10^{-14}

Critical Thinking Questions

- Are all the reactions in Table 1 that are equilibrium acids?
 (a) at equilibrium and yes
- Are all the reactions in Table 1 that are equilibrium bases?

Class work

- Are all the reactions in Table 1 that are Brønsted-Lowry acids?

Individual questions and answers

- Are all the reactions in Table 1 that are Brønsted-Lowry bases?

Group questions

- Are possible reactions that are not acid-base reactions?

Individual questions and answers

- Are all the reactions in Table 1 that are Brønsted-Lowry acids?

Class

- Are all the reactions in Table 1 that are Brønsted-Lowry bases?

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CH Cherryholmes



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