solution of the state of the second

#### Part D Questions and Problems

- An increase in the oxidation number of an atom indicates oxidation. A decrease in the oxidation number indicates reduction.
- 27. N is reduced (+5 to +2); Br is oxidized (-1 to 0) Mn is reduced (+7 to +2); Cl is oxidized (-1 to 0) N is reduced (+5 to +2); Sb is oxidized (0 to +5) S is reduced (+6 to +4); C is oxidized (0 to +4)

#### Section Review 20.3

#### Part A Completion

- L. coddation number
- 2. half-reaction
- two
   added
- 3. balanced

7. ionic

4. ionie

#### Part B True-False

- 8. AT 10. AT 9. AT 11. NT
- 12. NT 13. NT

200. (2)

#### Part C Matching

- 14. c 17. b 15. a 18. c
- 16. f 19. d

#### Part D Questions and Problems

a. HNO<sub>3</sub> → NO; N changes +5 to +2, a gain of 3e<sup>-</sup>; multiply by 2
 2H1 → 1<sub>2</sub>; I changes from −1 to 0, a loss of 2e<sup>-</sup> for I<sub>3</sub>; multiply by 3

2HNO<sub>3</sub> + 6HI → 2NO + 3I<sub>2</sub> + 4H<sub>2</sub>O

b. HNO<sub>3</sub> → NO<sub>3</sub>; N changes +5 to +4, a gain of 1e<sup>+</sup>; multiply by 10

 $I_2 \rightarrow 2HIO_3$ ; I changes from 0 to +5, a loss of  $10e^-$  for  $I_2$ 

 $10\mathrm{HNO_3} + \mathrm{I_2} \rightarrow 2\mathrm{HIO_3} + 10\mathrm{NO_2} + 4\mathrm{H_2O}$ 

22. a. S<sup>2-</sup> → S + 2e<sup>-</sup> and 3e<sup>-</sup> + 4H<sup>+</sup> + NO<sub>3</sub> → NO + 2H<sub>2</sub>O; Multiply the oxidation reaction by 3 and the reduction reaction by 2. 3H<sub>2</sub>S + 2HNO<sub>3</sub> → 3S + 2NO + 4H<sub>2</sub>O b. Fe<sup>2</sup> → Fe<sup>3</sup> + e<sup>-</sup> and 6e<sup>-</sup> + 14H<sup>+</sup> + Cr<sub>2</sub>O<sub>2</sub><sup>2</sup> → 2Cr<sup>3</sup> + 7H<sub>2</sub>O; multiply the oxidation reaction by 6 14H<sup>+</sup> + 6Fe<sup>2</sup> + Cr<sub>2</sub>O<sub>2</sub><sup>2</sup> → 6Fe<sup>2</sup> + 2Cr<sup>3</sup> + 7H<sub>2</sub>O

#### **Practice Problems**

#### Section 20.1

- Sr: oxidized (reducing agent)
- O<sub>2</sub>: reduced (oxidizing agent) 2. Li: oxidized (reducing agent)
- S: reduced (oxidizing agent)
- Cs: oxidized (reducing agent)
- Br<sub>3</sub>: reduced (oxidizing agent) 4. Mg: oxidized (reducing agent)
- N<sub>s</sub>: reduced (oxidizing agent)
- 5. Fe: oxidized (reducing agent)
  - O<sub>2</sub>: reduced (oxidizing agent)
- Br": oxidized (reducing agent)
- Cl<sub>2</sub>: reduced (oxidizing agent)
- Si: oxidized (reducing agent)
   F<sub>2</sub> reduced (oxidizing agent)
- Ca: exidized (reducing agent)
   O<sub>2</sub> reduced (exidizing agent)
- O<sub>2</sub> reduced (oxidizing agent)
  9. Mg: oxidized (reducing agent)
  - H': reduced (oxidizing agent)
- 10. Na: oxidized (reducing agent)
- H<sub>2</sub>O: reduced (exidizing agent)

#### Section 20.2

- a. Sn is tin in an uncombined state. The oxidation number is 0.
  - b. The ionic charge on potassium is 1+, thus the oxidation number is +1.
  - The ionic charge on sulfur is 2-, thus the oxidation number is -2.
  - d. The ionic charge on iron is 3+, thus the oxidation number is +3.
  - e. Se is selenium in an uncombined state. The oxidation number is 0
  - f. The ionic charge on magnesium is 2+, thus the exidation number is +2.
  - g. The ionic charge on tin is 4+, thus the oxidation number is +4.

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