

CHAPTER 3 TEST*Atoms: The Building Blocks of Matter*

MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question.

1. _____ The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays were composed of _____.
- a. energy
 - b. positive particles
 - c. negative particles
 - d. neutral particles
2. _____ The basic principles of atomic theory that are still recognized today were first conceived by _____.
- a. Avogadro
 - b. Bohr
 - c. Dalton
 - d. Rutherford
3. _____ An example of the law of multiple proportions is the existence of _____.
- a. FeCl_2 and $\text{Fe(NO}_3)_3$
 - b. O_2 and O_3
 - c. CO and CO_2
 - d. FeCl_2 and $\text{Fe(NO}_3)_2$
4. _____ Atoms of the same element can differ in _____.
- a. chemical properties
 - b. mass number
 - c. atomic number
 - d. number of protons and electrons
5. _____ Dalton's atomic theory helped to explain the law of conservation of mass because it stated that atoms _____.
- a. could not combine
 - b. were indivisible
 - c. all had the same mass
 - d. could not be created or destroyed
6. _____ Millikan's experiments determined _____.
- a. that the electron carried no charge
 - b. that the electron carried the smallest possible positive charge
 - c. the approximate value of the electron's mass
 - d. that the electron had no mass
7. _____ In Rutherford's experiment, a small percentage of the positively charged particles bombarding the metal's surface _____.
- a. were slightly deflected as they passed through the metal
 - b. were deflected back toward the source from the metal
 - c. passed straight through the metal
 - d. combined with the metal
8. _____ Most of the volume of an atom is made up of the _____.
- a. nucleus
 - b. nucleon
 - c. electron cloud
 - d. proton

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