

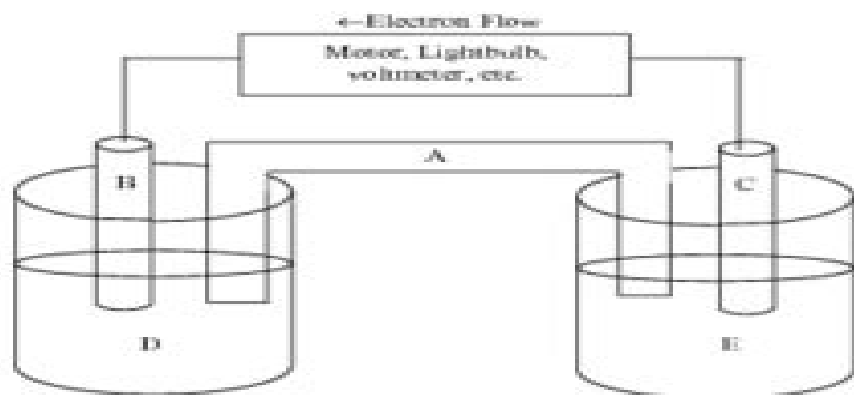
Honors Chemistry

Electrochemistry Problems

Name _____

Date _____ Period _____

The voltaic cell shown above is composed of the metals Nickel (Ni) and Calcium (Ca) and their nitrate salts. Assume both metals are 2^+ in their ionic state. Pay attention to the direction of electron flow. Answer the following questions using the letters indicated above. The same letter may be used more than once.



1. Salt bridge
2. Nickel electrode
3. Calcium nitrate solution
4. Calcium electrode
5. Nickel nitrate solution
6. Anode
7. Cathode
8. Electrode where oxidation occurs
9. Electrode where reduction occurs
10. KNO_3
11. Electrode that gets smaller with time
12. Solution that receives positive ions from the salt bridge
13. If the reduction potentials for Ni and Ca are -0.25V and -2.87V respectively, what is the voltage of the cell?

Write a condensed cell diagram for this reaction.

Write the oxidation half reaction for the cell.

Write the reduction half reaction for the cell.

For the following five questions identify the half reactions as either oxidized or reduced

14. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + 1e^-$
15. $\text{Cu} \rightarrow \text{Cu}^{2+} + 1e^-$
16. $\text{Cl}_2 + 2e^- \rightarrow 2\text{Cl}^-$
17. $\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$
18. $2\text{N}^{3-} \rightarrow \text{N}_2 + 6e^-$
19. For the reaction

$$2\text{Mn}^{2+} + 3\text{Cl}_2 \rightarrow 2\text{Mn}^{3+} + 6\text{Cl}^-$$
 which species is the reducing agent? Explain.

High School Electrochemistry Questions

Daniela Niemeyer



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Table of Contents High School Electrochemistry Questions

1. Understanding the eBook High School Electrochemistry Questions
 - The Rise of Digital Reading High School Electrochemistry Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying High School Electrochemistry Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an High School Electrochemistry Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from High School Electrochemistry Questions
 - Personalized Recommendations
 - High School Electrochemistry Questions User Reviews and Ratings
 - High School Electrochemistry Questions and Bestseller Lists
5. Accessing High School Electrochemistry Questions Free and Paid eBooks

- High School Electrochemistry Questions Public Domain eBooks
- High School Electrochemistry Questions eBook Subscription Services
- High School Electrochemistry Questions Budget-Friendly Options
- 6. Navigating High School Electrochemistry Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - High School Electrochemistry Questions Compatibility with Devices
 - High School Electrochemistry Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of High School Electrochemistry Questions
 - Highlighting and Note-Taking High School Electrochemistry Questions
 - Interactive Elements High School Electrochemistry Questions
- 8. Staying Engaged with High School Electrochemistry Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers High School Electrochemistry Questions
- 9. Balancing eBooks and Physical Books High School Electrochemistry Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection High School Electrochemistry Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine High School Electrochemistry Questions
 - Setting Reading Goals High School Electrochemistry Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of High School Electrochemistry Questions
 - Fact-Checking eBook Content of High School Electrochemistry Questions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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