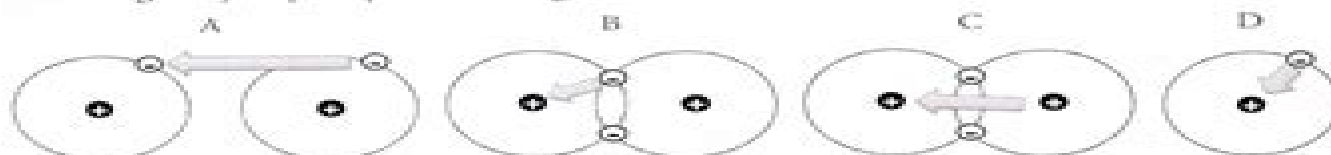


Read This!

Electronegativity is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

Electronegativity decreases as you go down a group (vertical column).

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.

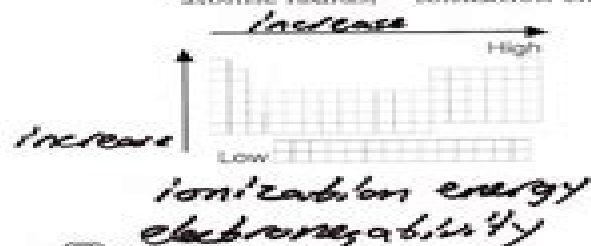
- c. What is the trend in electronegativity going across a period in Model 1?

Electronegativity increases from left to right across a period (horizontal row)

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



High Marks Chemistry Answer Key

Henry Lester Gerry



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 - The Rise of Digital Reading High Marks Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying High Marks Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an High Marks Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from High Marks Chemistry Answer Key
 - Personalized Recommendations
 - High Marks Chemistry Answer Key User Reviews and Ratings
 - High Marks Chemistry Answer Key and Bestseller Lists
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 - High Marks Chemistry Answer Key Budget-Friendly Options
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 - High Marks Chemistry Answer Key Enhanced eBook Features
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- Adjustable Fonts and Text Sizes of High Marks Chemistry Answer Key
 - Highlighting and Note-Taking High Marks Chemistry Answer Key
 - Interactive Elements High Marks Chemistry Answer Key
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers High Marks Chemistry Answer Key
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 - Creating a Diverse Reading Collection High Marks Chemistry Answer Key
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 - Minimizing Distractions
 - Managing Screen Time
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- Setting Reading Goals High Marks Chemistry Answer Key
 - Carving Out Dedicated Reading Time
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- Fact-Checking eBook Content of High Marks Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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