

## Mole Calculation Practice Worksheet

Answer the following questions:

- 1) How many moles are in 25 grams of water?

$$\frac{25\text{g H}_2\text{O}}{18.02\text{g H}_2\text{O}} \times \frac{1\text{ mol H}_2\text{O}}{1\text{ mol H}_2\text{O}} = 1.4 \text{ moles}$$

- 2) How many grams are in 4.5 moles of  $\text{Li}_2\text{O}$ ?

$$\frac{4.5\text{ mol Li}_2\text{O}}{1\text{ mol Li}_2\text{O}} \times \frac{29.88\text{g Li}_2\text{O}}{1\text{ mol Li}_2\text{O}} = 130 \text{ g}$$

- 3) How many molecules are in 23 moles of oxygen?

$$\frac{23\text{ mol O}_2}{1\text{ mol O}_2} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1\text{ mol O}_2} = 1.4 \times 10^{25} \text{ molecules}$$

- 4) How many moles are in  $3.4 \times 10^{23}$  molecules of  $\text{H}_2\text{SO}_4$ ?

$$\frac{3.4 \times 10^{23} \text{ molecules H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \times \frac{1\text{ mol H}_2\text{SO}_4}{1\text{ mol H}_2\text{SO}_4} = 0.56 \text{ moles}$$

- 5) How many molecules are in 25 grams of  $\text{NH}_3$ ?

$$\frac{25\text{g NH}_3}{17.03\text{g NH}_3} \times \frac{1\text{ mol NH}_3}{1\text{ mol NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules NH}_3}{1\text{ mol NH}_3} = 8.8 \times 10^{23} \text{ molecules}$$

- 6) How many grams are in  $8.2 \times 10^{22}$  molecules of  $\text{N}_2\text{I}_4$ ?

$$\frac{8.2 \times 10^{22} \text{ molecules N}_2\text{I}_4}{6.02 \times 10^{23} \text{ molecules N}_2\text{I}_4} \times \frac{1\text{ mol N}_2\text{I}_4}{1\text{ mol N}_2\text{I}_4} \times \frac{789.42\text{g N}_2\text{I}_4}{1\text{ mol N}_2\text{I}_4} = 110 \text{ g}$$

# Grams Moles Calculations Worksheet Answer Key

**Michal Rosen-Zvi**

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