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Chapter 2: Atoms

- 2.1** (a) NaClO_3 (b) AlF_3
- 2.2** (a) The mass number is $15 + 16 = 31$.
(b) The mass number is $86 + 136 = 222$.
- 2.3** (a) The element has 15 protons, making it phosphorus (P); its symbol is ${}_{15}^{31}\text{P}$.
(b) The element has 86 protons, making it radon (Rn); its symbol is ${}_{86}^{222}\text{Rn}$.
- 2.4** (a) The atomic number of mercury (Hg) is 80; that of lead (Pb) is 82.
(b) An atom of Hg has 80 protons; an atom of Pb has 82 protons.
(c) The mass number of this isotope of Hg is $80 + 120 = 200$; the mass number for this isotope of Pb is $82 + 120 = 202$.
(d) The symbols of these isotopes are ${}^{200}_{80}\text{Hg}$ and ${}^{202}_{82}\text{Pb}$.
- 2.5** The atomic number of iodine (I) is 53. The number of neutrons in each isotope is $125 - 53 = 72$ for iodine-125 and $131 - 53 = 78$ for iodine-131. The symbols for these two isotopes are ${}^{125}_{53}\text{I}$ and ${}^{131}_{53}\text{I}$.
- 2.6** The atomic weight is 6.941 amu, which is nearer to 7 amu than 6 amu. Therefore, lithium-7 is the more abundant isotope. The relative abundances for these two isotopes are 92.50 percent for lithium-7 and 7.50 percent for lithium-6.
- 2.7** This element has 13 electrons and, therefore, 13 protons. The element with atomic number 13 is Aluminum (Al).
- Ât :**
- 2.8** Both Democritus and Dalton believed that matter was composed of tiny indivisible particles referred to as atoms. The major difference between Democritus and Dalton is that Dalton based his theory on evidence rather than belief.
- 2.10** (a) Oxygen - an element (b) Table salt - a compound
(c) Sea water - a mixture (d) Wine - a mixture
(e) Air - a mixture (f) Silver - an element
(g) Diamond - an element (h) A pebble - a mixture
(i) Gasoline - a mixture (j) Milk - a mixture
(k) Carbon dioxide - a compound (l) Bronze - a mixture

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Table of Contents General Organic And Biochemistry Answers To Test

1. Understanding the eBook General Organic And Biochemistry Answers To Test
 - The Rise of Digital Reading General Organic And Biochemistry Answers To Test
 - Advantages of eBooks Over Traditional Books
2. Identifying General Organic And Biochemistry Answers To Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Organic And Biochemistry Answers To Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Organic And Biochemistry Answers To Test
 - Personalized Recommendations
 - General Organic And Biochemistry Answers To Test User Reviews and Ratings
 - General Organic And Biochemistry Answers To Test and Bestseller Lists
5. Accessing General Organic And Biochemistry Answers To Test Free and Paid eBooks
 - General Organic And Biochemistry Answers To Test Public Domain eBooks
 - General Organic And Biochemistry Answers To Test eBook Subscription Services
 - General Organic And Biochemistry Answers To Test Budget-Friendly Options
6. Navigating General Organic And Biochemistry Answers To Test eBook Formats
 - ePub, PDF, MOBI, and More
 - General Organic And Biochemistry Answers To Test Compatibility with Devices
 - General Organic And Biochemistry Answers To Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Organic And Biochemistry Answers To Test
 - Highlighting and Note-Taking General Organic And Biochemistry Answers To Test
 - Interactive Elements General Organic And Biochemistry Answers To Test

8. Staying Engaged with General Organic And Biochemistry Answers To Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Organic And Biochemistry Answers To Test
9. Balancing eBooks and Physical Books General Organic And Biochemistry Answers To Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Organic And Biochemistry Answers To Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Organic And Biochemistry Answers To Test
 - Setting Reading Goals General Organic And Biochemistry Answers To Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Organic And Biochemistry Answers To Test
 - Fact-Checking eBook Content of General Organic And Biochemistry Answers To Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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