

# INTERMOLECULAR FORCES – ANSWER KEY (from Tro, Chapter 11, page 512)

49. Determine the kinds of intermolecular forces that are present in each of the following elements or compounds:

- |                      |                          |                        |                        |
|----------------------|--------------------------|------------------------|------------------------|
| (a) Kr*              | (b) NCl <sub>3</sub> **  | (c) SiH <sub>4</sub> * | (d) HF****             |
| (e) N <sub>2</sub> * | (f) NH <sub>3</sub> **** | (g) CO**               | (h) CCl <sub>4</sub> * |

50. Determine the kinds of intermolecular forces that are present in each of the following elements or compounds:

- |                        |                          |                            |                      |
|------------------------|--------------------------|----------------------------|----------------------|
| (a) HCl**              | (b) H <sub>2</sub> O**** | (c) Br <sub>2</sub> *      | (d) He*              |
| (e) PH <sub>3</sub> ** | (f) HBr**                | (g) CH <sub>3</sub> OH**** | (h) I <sub>2</sub> * |

\* dispersion forces only; \*\* dispersion and dipole-dipole forces; \*\*\*\* dispersion force, dipole-dipole forces, and hydrogen bonds

51. Arrange the following in order of increasing boiling point. Explain your reasoning.

- |                     |                                     |  |   |
|---------------------|-------------------------------------|--|---|
| (a) CH <sub>4</sub> | (b) CH <sub>3</sub> CH <sub>3</sub> | (c) CH <sub>3</sub> CH <sub>2</sub> Cl | (d) CH <sub>3</sub> CH <sub>2</sub> OH <b>a b c d</b> |
|---------------------|-------------------------------------|--|---|

52. Arrange the following in order of increasing boiling point. Explain your reasoning.

- |                      |                       |                                   |
|----------------------|-----------------------|-----------------------------------|
| (a) H <sub>2</sub> S | (b) H <sub>2</sub> Se | (c) H <sub>2</sub> O <b>a b c</b> |
|----------------------|-----------------------|-----------------------------------|

53. For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning.

- |  |   |
|--|---|
| (a) <b>CH<sub>3</sub>OH</b> or CH <sub>3</sub> SH          | (b) CH <sub>3</sub> OCH <sub>3</sub> or <b>CH<sub>3</sub>CH<sub>2</sub>OH</b> |
| (c) CH <sub>4</sub> or <b>CH<sub>3</sub>CH<sub>3</sub></b> |   |

54. For each pair of compounds, pick the one with the higher boiling point. Explain your reasoning.

- |  |  |  |
|--|--|--|
| (a) <b>NH<sub>3</sub></b> or CH <sub>4</sub> | (b) <b>CS<sub>2</sub></b> or CO <sub>2</sub> | (c) CO <sub>2</sub> or <b>NO<sub>2</sub></b> |
|--|--|--|

55. For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning.

- |   |   |  |
|---|---|--|
| (a) <b>Br<sub>2</sub></b> or I <sub>2</sub> | (b) <b>H<sub>2</sub>S</b> or H <sub>2</sub> O | (c) NH <sub>3</sub> or <b>PH<sub>3</sub></b> |
|---|---|--|

56. For each pair of compounds, pick the one with the higher vapor pressure at a given temperature. Explain your reasoning.

- |   |
|---|
| (a) <b>CH<sub>4</sub></b> or CH <sub>3</sub> Cl                                   |
| (b) CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> OH or <b>CH<sub>3</sub>OH</b> |
| (c) CH <sub>3</sub> OH or <b>H<sub>2</sub>CO</b>                                  |

57. Which of the following pairs of substances would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved.

- |   |  |
|---|--|
| (a) CCl <sub>4</sub> and H <sub>2</sub> O                         | (b) <b>KCl and H<sub>2</sub>O (ion-dipole forces)</b>                  |
| (c) <b>Br<sub>2</sub> and CCl<sub>4</sub> (dispersion forces)</b> | (d) <b>CH<sub>3</sub>CH<sub>2</sub>OH and H<sub>2</sub>O (H-bonds)</b> |

58. Which of the following pairs of compounds would you expect to form solutions (homogeneous mixtures) when combined? For those that form solutions, indicate the type of forces that are involved.

- |  |
|--|
| (a) <b>CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub> and CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub> (dispersion forces)</b> |
| (b) CBr <sub>4</sub> and H <sub>2</sub> O  |
| (c) <b>LiNO<sub>3</sub> and H<sub>2</sub>O (ion-dipole forces)</b>   |
| (d) CH <sub>3</sub> OH and CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>   |

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**A Gutmann**



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