

# Honors Calorimetry

*Instructions: Complete each section of this report and answer all questions in complete sentences for full credit.*

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**Materials:** thermometer, two foam cups, measuring cup from your kitchen, hot water, and one liquid selected from the chart below and placed in your refrigerator for one or more hours before you begin your activity

Liquid	Heat capacity (J/(g °C))	Density (g/mL)
Grape juice	3.6	1.06
Whole milk	3.77	1.04
Orange juice	3.73	1.05
Tomato juice	3.98	1.03

**Caution:** You will be working with hot water in this lab. Be very careful as you pour so you do not burn yourself. Also, be sure to read your thermometer carefully. The numbers on the thermometer scale refer to the line under the number. Lastly, it's important to perform certain steps quickly and efficiently to minimize heat loss to the surroundings.

## Procedure:

1. Read the entire procedure before you begin the experiment.
2. Use the metric scale on your measuring cup to measure 237 mL (1 cup) of your selected refrigerated liquid. Pour this liquid into one of the foam cups.
3. Now, get the hottest water possible from the faucet on your sink. Use the metric scale on your measuring cup to measure 237 mL (1 cup) of hot water. Pour this water into the other foam cup.
4. Quickly measure the temperature of the cold liquid and the hot water. Record the temperatures in a data table.
5. Very quickly pour the hot water into the cold liquid. Place the thermometer in the mixture. Invert the cup that contained the cold liquid over the cup containing the water mixture.
6. Keeping the bulb of the thermometer in the liquid, check the temperature reading about every 20 seconds until the temperature stabilizes. Keep the lid (inverted cup) in place as much as possible. When the temperature is constant (this will happen very quickly), record this temperature in a data table as the final temperature.

## Data

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# Flvs Chemistry Honors Answers

**Lingsheng Yao**



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