

# 10th Grade Honors Chemistry Final Exam Flashcards With Complete Solutions

**Thermochemistry CORRECT ANSWERS** The study of heat effects in chemical reactions.

**Exothermic Reaction CORRECT ANSWERS** A reaction that releases energy in the form of heat.

Examples: Combustion, freezing water.

Heat generated is added to the product side (right side) because heat is released.

**Endothermic Reaction CORRECT ANSWERS** A reaction that absorbs energy in the form of heat.

Examples: Melting ice, evaporating liquid water, boiling water.

Heat generated is added to the reactant side (left side) because heat is added.

**Enthalpy (H or  $\Delta H$ ) CORRECT ANSWERS** The energy after taking into account the pressure and volume of a substance.

If pressure remains constant, then heat absorbed or released during a chemical reaction is equal to the enthalpy change for the reaction.

$\Delta H =$  CORRECT ANSWERS  $H_{\text{Products}} - H_{\text{Reactants}}$ .

**Standard State of Element (Temperature and Pressure) CORRECT ANSWERS**  
Temperature = 25° Celsius.  
Pressure = 1 atm.

$\Delta H^\circ$  CORRECT ANSWERS Standard enthalpy change = the heat transferred in a reaction with reactants and products in standard states.

**Determining Enthalpy Change CORRECT ANSWERS**  $\Delta H = (-)$  Exo.  
 $\Delta H = (+)$  Endo.

If heat is released as in an exothermic process the temperature will increase, but the product will have less energy than the reactant so enthalpy will be negative.

# Honors Chemistry Final Exam Mc Practice

**Lei Huang**



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