## Gas Stoichiometry: Multiple Choice

 What volume of pure oxygen gas, O<sub>2</sub>, measured at 546 K and 101.325 kPa is formed by the complete decomposition of 0,200 mol Ag<sub>2</sub>O according to the following equation:

2Ag/O → 4Ag + O<sub>2</sub>

A) 1.12 L B) 4.48 L C) 2.24 L D) 8.96 L E) 17.82 L

ANS: B

I. How many moles of KClO<sub>2</sub> are needed to form 2.8 L of O<sub>2</sub>, measured at STP, according to the following reaction:  $2KClO_2 \rightarrow 2KCl + 3O_2$ 

A) 0.083 mol B) 0.25 mol C) 0.17 mol D) 0.50 mol E) 0.33 mol

ANS: A

1. What volume of chlorine gas, measured at 0°C and 101,325 kPa, is required to react completely with 5.61 g of KOH in the following reaction:  $Cl_{2m} + 2KOH \rightarrow KCI + KCIO + H_{2m}$ 

A) 0.56 L B) 4.48 L C) 1.12 L D) 22.4 L E) 2.24 L

ANS: C

Sodium hydride reacts with excess water to produce sodium hydroxide (aq) and hydrogen gas: NaH(s)
H (20) = NaCH(sex) + H (a)

+ H<sub>2</sub>O(l) → NaOH(aq) + H<sub>2</sub>(g)

A sample of Nall weighing \_\_\_\_\_ g will produce 982 mL of gas at 28.0 °C and 765 torr, when the hydrogen is collected over water. The vapor pressure of water at this temperature is 28 torr.

A) 2.93 B) 0.960 C) 0.925 D) 0.0388 E) 925

ANS: C

1. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to

form water?  $2H^{2}(g) + O^{2}(g) \rightarrow 2H^{2}O(g)$ 

A) 1.8 L B) 3.6 L C) 2.0 L D) 2.4 L

ANS: A

The molar volume of a gas at STP is \_\_\_\_\_\_ L.

A) 0.08206 B) 62.36 C) 1.00 D) 22.4 E) 14.7

ANS: D

 If 0.5 L of O<sub>2</sub>(g) reacts with H<sub>2</sub> to produce 1 L of H<sub>2</sub>O(g), what is the volume of H<sub>2</sub>O(g) obtained from 1 L of O<sub>3xi</sub>?

A) 0.5 L B) 1.5 L C) 2 L D) 2.5 L

ANS: C

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