1)Which of the following statements is incorrect?

- A) The degree of ionization of a weak electrolyte increases with dilution.
- B) Strong electrolytes are ionized completely at all concentrations.
- C) Addition of NH₄CL to NH₄OH increases the ionization of the latter.
- D) Increase the temperature increases the ionization.

2) According to Lowry and Bronsted system the chloride ion (CI) in aqueous solution is a

- A) Weak base
- B) Strong base
- C) Weak acid
- D) Strong acid

3) In the following reactions, ZnO is respectively acting as a/an

¡)ZnO+Na₂O gives Na₂ZnO₂

- (i) ZnO+CO, gives ZnCO,
 - A) Acid and acid
 - B) Acid and base
 - C) Base and acid
 - D) Base and base

4) Given $k_{a1}k_{a2}$ for $H_2S=1.1\times10^{-21}M^3$. 1L of 0.1 M H_2S having [H*]=0.1M contains S^{2^4} ions equal to.

- A) 6.62×103
- B) 6.62×104
- C) 6.62×10°
- D) 6.62×10°

5) Equal volume of three acid solutions of pH 3.4 and 5 are mixed in a vessel. What will be

Multiple Choice Questions On Chemical Equilibrium

Tobias Bleicker

Multiple Choice Questions On Chemical Equilibrium:

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Reviewing Multiple Choice Questions On Chemical Equilibrium: Unlocking the Spellbinding Force of Linguistics

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