#### Multiple Choice Questions

- K<sub>1</sub> and K<sub>2</sub> are the velocity constants of forward and backward reactions.
  The equilibrium constant k of the reaction is
  - (A)  $K_j \times K_j$

(B)  $K_1 - K_2$ 

(C)  $\frac{K_I}{K_I}$ 

 $(D) \frac{K_I + K_I}{K_I - K_I}$ 

- Sed. (C).
- Which of the following reactions will be favoured at low pressure?
  - (A) N, +0, -2NO

- (B) H<sub>2</sub> + I<sub>2</sub> → 2HI
- (C)  $PCI_A \longrightarrow 2PCI_A + CI_A$
- (D)  $N_j + 3H_j = 5NH_j$

- Sed. (C).
- For the reaction, A+B--+C+D, if active mass of A is kept constant and active mass of B is tripled, the rate of reaction will become
  - (A) three times

(B) six times

(C) eight times

(D) nine times

- Sol. (D).
- 4. The rate of  $\frac{K_p}{K_s}$  for the reaction:

$$CO(g) + \frac{1}{2}O_{1}(g) = CO_{1}(g)$$
 is

(A), I,

(B) RT

 $(C) (RT)^{I/I}$ 

(D)  $(RT)^{-\epsilon}$ 

- Sol. (D).
- 5. Consider the reaction,

in a closed container at equilibrium. What would be the effect of addition of CaCO, on the equilibrium concentration of CO2?

(A) Increases

tBY Unpredictable

(C) Decreases:

(D) Remains unaffected

- Sed. (106.)
- For the reaction, PG<sub>g</sub>(g), PG<sub>g</sub>(g) + G<sub>g</sub>(g), the forward reaction at constant temperature is favoured by
  - (A) introducing an mert gas at constant volume
  - (B) introducing PCI(g) at constant volume
  - (C) introducing PCl-(g) at constant volume
  - (D) introducing Cl3(g) at constant volume
- Sol. (C).
- A liquid is in equilibrium with its vapour at its boiling point. On an average, the molecules in the two phases have equal
  - (A) intermolecules forces
  - (B) potential energy
  - (C) kimeric energy
  - (D) both kinetic energy and potential energy
- Soul. (C).
- 8. The unit of equilibrium constant for the reversible reaction

# **Multiple Choice Questions In Chemical Equilibrium**

**Jin-Ying Zhang** 

Multiple Choice Questions In Chemical Equilibrium:

#### Multiple Choice Questions In Chemical Equilibrium Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has be much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Multiple Choice Questions In Chemical Equilibrium**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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