

2.1. Thermodynamics (pp. 12-15)

Thermodynamics is the study of energy and its transformations. The first law of thermodynamics states that energy cannot be created or destroyed, only transformed from one form to another.

1. Internal Energy is the total energy contained within a system. It is the sum of the kinetic and potential energies of all the particles in the system.

2. The change in internal energy of a system is given by:

$\Delta U = Q - W$

where  $\Delta U$  is the change in internal energy,  $Q$  is the heat added to the system, and  $W$  is the work done by the system.

3. Heat is energy transferred between two systems due to a temperature difference.

4. Work is energy transferred between two systems by a force acting through a distance.

5. The first law of thermodynamics can be written as:  $\Delta U = Q - W$

6. In a thermodynamic process, the internal energy of a system can change due to heat transfer and work done.

7. The heat added to a system is  $Q$  and the work done by the system is  $W$ .

8. The change in internal energy of a system is  $\Delta U$ .

9. The first law of thermodynamics states that the total energy of an isolated system is constant.

10. The internal energy of a system is a state function, meaning it depends only on the state of the system and not on the path taken to reach that state.

11. The heat added to a system is  $Q$  and the work done by the system is  $W$ .

12. The change in internal energy of a system is  $\Delta U$ .

13. The first law of thermodynamics can be written as:  $\Delta U = Q - W$

14. The heat added to a system is  $Q$  and the work done by the system is  $W$ .

15. The change in internal energy of a system is  $\Delta U$ .

# Homeostasis Directed Reading Guide Answers

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