Hess's Law

- Hess's law of constant heat summation
 - If a process occurs in stages or steps (even hypothetically),
 the enthalpy change for the overall process is the sum of the enthalpy changes for the individual steps.

$$^{1/2}N_{2}(g) + ^{1/2}O_{2}(g) \rightarrow NO(g)$$
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 $NO(g) + ^{1/2}O_{2}(g) \rightarrow NO_{2}(g)$ $\Delta H = -57.07 \text{ kJ}$

$$\frac{1}{2}N_{2}(g) + O_{2}(g) \rightarrow NO_{2}(g)$$
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