

General Chemistry Review (Chapters 1 – 3)

1. The first two rows of the periodic table, excluding noble gases and including larger halogens, is sufficient for most organic chemistry applications. Try to fill in the periodic table below from memory, then circle the most electronegative element.

1	2	3	4	5	6	7
1						
3	4	5	6	7	8	9
11	12	13	14	15	16	17
						35
						53

2. Rank the following atoms from largest to smallest: S, O, Br, I
3. How many and what type of orbitals are in the second electron shell? Draw the shape and orientation of each.
4. Identify the group of elements that corresponds to the following generalized electron configuration: [noble gas] ns^2np^3
5. Draw Lewis structures for the following molecular formulas. Don't forget lone pairs. (Hint: these molecules contain only SINGLE bonds.)



6. Draw Lewis structures for the following molecular formulas. Include lone pairs. (Hint: these molecules contain double and/or triple bonds.)



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William Collins



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