

## Answer Key

### Chapter 10

#### Practice Problems

1.  $9.45 \times 10^{24}$  molecules  $\text{CO}_2$
2.  $3.27 \times 10^{22}$  molecules  $\text{H}_2\text{O}$
3. 1.53 mol Fe
4. 0.0341 mol sucrose
5.  $1.08 \times 10^{21}$  formula units  $\text{CuSO}_4$
6. 839 g Sb
7. 5.53 g Se
8. 6.95 mol S
9. 12.5 mol He
10. 0.0213 mol Ni
11.  $6.17 \times 10^{21}$  atoms Pt
12.  $1.88 \times 10^{28}$  atoms S
13. 39.7 g Hg
14.  $6.35 \times 10^{-3}$  g
15.  $4.36 \times 10^{16}$  Pb ions
16. 0.220 mol  $\text{C}_6\text{H}_6$
17. 2.856 mol  $\text{KClO}_3$
18. 40.2 g  $\text{SnSO}_4$
19.  $3.40 \times 10^2$  g (340 g expressed to 3 significant figures)
20. 13.1 g Zn and 37.5 g  $\text{Cu}(\text{NO}_3)_2$
21. 52.93% Al, 47.07% O
22. 16.39% Mg, 18.89% N, 64.72% O
23. 39.17% O
24. 49.30% N
25. The percent compositions will be the same because the ratio of carbon atoms to hydrogen atoms is the same, 1:1.
26.  $\text{CH}_2\text{O}$
27.  $\text{ZnN}_2\text{O}_6$
28.  $\text{Al}_4\text{C}_3$
29.  $\text{C}_3\text{H}_4\text{O}_3$
30.  $\text{C}_8\text{H}_8\text{N}_2\text{O}_2$
31.  $\text{B}_3\text{H}_6\text{N}_3$
32.  $\text{Ag}_2\text{C}_2\text{O}_4$
33.  $\text{P}_2\text{S}_5$ ; same as empirical formula
34.  $\text{C}_9\text{H}_{12}\text{N}_6$
35.  $\text{CeI}_3 \cdot 9\text{H}_2\text{O}$
36.  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
37.  $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$
38.  $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$

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