

Wintersemester 2019/2020 - Klausuraufgaben 8B

Aufgabe 8 (max. 70 Minuten/100 Punkte)

Die nachstehenden Probleme betreffen die Veranschaulichung des Verhaltens der Stoffeigenschaften mittels "Diagramm" von Molwärme, Enthalpie, Entropie, etc. für ein bestimmtes gas. Bitte führen Sie alle nötigen Schritte zur Veranschaulichung strikt an! (Jeweils 20 Punkte)

$$n(P_1, T_1) = n(P_2, T_2)$$

1. Ein bestimmtes einatomiges Gas wird bei einer konstanten RTW erhitzt und bei einer konstanten Volumen, also bei einem konstanten Molvolumen, abgekühlt.

$$P_1 = 1 \text{ bar, } T_1 = 300 \text{ K, } P_2 = 10 \text{ bar, } T_2 = 100 \text{ K}$$

2. Ein bestimmtes Gas wird bei einer konstanten RTW bei 10 bar erhitzt und bei einem konstanten Molvolumen abgekühlt.

$$P_1 = 10 \text{ bar, } T_1 = 100 \text{ K, } P_2 = 1 \text{ bar, } T_2 = 300 \text{ K}$$

3. Ein bestimmtes Gas wird bei einer konstanten RTW bei 10 bar erhitzt und bei einer konstanten Molvolumen abgekühlt. Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie?

$$P_1 = 10 \text{ bar, } T_1 = 100 \text{ K, } P_2 = 1 \text{ bar, } T_2 = 200 \text{ K}$$

4. Ein bestimmtes Gas wird bei einer konstanten RTW bei 10 bar erhitzt und bei einer konstanten Molvolumen abgekühlt. Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie? Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie? Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie?

$$P_1 = 10 \text{ bar, } T_1 = 100 \text{ K, } P_2 = 1 \text{ bar, } T_2 = 200 \text{ K}$$

5. Ein bestimmtes Gas wird bei einer konstanten RTW bei 10 bar erhitzt und bei einer konstanten Molvolumen abgekühlt. Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie? Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie?

$$P_1 = 10 \text{ bar, } T_1 = 100 \text{ K, } P_2 = 1 \text{ bar, } T_2 = 200 \text{ K}$$

6. Ein bestimmtes Gas wird bei einer konstanten RTW bei 10 bar erhitzt und bei einer konstanten Molvolumen abgekühlt. Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie? Die Temperaturerhöhung beträgt 100 K und die Temperaturerniedrigung 100 K. Wie groß ist die Änderung der Enthalpie?

Glenco Contentent Mastery Section Behavior Of Gases Answers

Clemens Wendtner



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Unveiling the Magic of Words: A Review of "**Glenco Contentent Mastery Section Behavior Of Gases Answers**"

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