

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- copper (II) oxide CuO
- mercury (I) oxide Hg₂O
- gold (III) chloride AuCl₃
- thallium (III) bromide TlBr₃
- bismuth (V) oxide Bi₂O₅
- terbium (III) oxide Tb₂O₃
- uranium (VI) oxide UO₃
- protactinium (IV) bromide PaBr₄
- cerium (III) oxide Ce₂O₃
- arsenic (V) sulphide As₂S₅

- manganese (III) oxide Mn₂O₃
- vanadium (II) bromide VBr₂
- niobium(V) oxide Nb₂O₅
- titanium (III) oxide Ti₂O₃
- titanium (III) nitride TiN
- iron (II) oxide FeO
- cobalt (II) phosphide Co₃P₂
- tin (II) oxide SnO
- thulium (II) bromide TmBr₂
- copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- SnCl₄ Tin (IV) Chloride
- BiBr₅ Bismuth (V) Bromide
- PoO₂ Polonium (IV) Oxide
- PbI₂ Lead (II) Iodide
- HgO Mercury (II) Oxide
- HgCl Mercury (I) Chloride
- Au₂O₃ Gold (III) Oxide
- FeCl₂ Iron (II) Chloride

- PdF₄ Palladium (IV) Fluoride
- Os₂O₃ Osmium (III) Oxide
- MoBr₂ Molybdenum (II) Bromide
- VCl₆ Vanadium (V) Chloride
- Mn₂O₃ Manganese (III) Oxide
- CoO Cobalt (II) Oxide
- Np₂O₃ Neptunium (III) Oxide
- V₂O₅ Vanadium (V) Oxide

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