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# **ECONOMICS PAPER 1**

**NATIONAL  
SENIOR CERTIFICATE**

**SECTION A 2023**

**GRADE 12**

**FINAL REVISION**



- 1. The graph shows the relationship between price and quantity.
- 2. The upward-sloping line represents the supply curve.
- 3. The downward-sloping line represents the demand curve.
- 4. The intersection of the two curves is the equilibrium point.
- 5. The equilibrium price is the price at which the quantity demanded equals the quantity supplied.
- 6. The equilibrium quantity is the quantity at which the quantity demanded equals the quantity supplied.

# Final Examination Scope For Grade 1economics Paper 2

**WJ Hussar**



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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds  $\text{Cl}^-\text{aq}$

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver  $\text{I}^+$  ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4  $13\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$  because two  $\text{NH}_4^+\text{aq}$

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