

# Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

## 1. Physical Science Basics

- What tool(s) would you want to use to find the:
  - Mass of an object? triple beam balance Basic SI Unit? kg
  - Volume of an object? ruler or water displacement method Basic SI Unit? cm<sup>3</sup> or ml
  - Length of an object? Ruler Basic SI Unit? meter
  - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
  - What is the density of a cube that is 25 g and has a volume of 3 cm<sup>3</sup>? 8g/cm<sup>3</sup>
  - Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has. solid H<sub>2</sub>O less dense than liquid H<sub>2</sub>O

## 2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

## 3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1802	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell ( <u>energy level</u> ) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

## 4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
  - An element's properties can be predicted from its location in the periodic table.
  - The atomic number represents the number of protons.
  - The atomic mass represents the number of neutrons and protons.

# Final Exam Study Guide Physical Science

**H. Lee Swanson**



## Final Exam Study Guide Physical Science:

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procedures and article characteristics Addressing such questions as the merits of inclusion settings and the relative benefits of direct and strategy instruction Swanson offers timely recommendations for instructional design assessment and policy

*Collateral Damage* Sharon L. Nichols, David C. Berliner, 2007-03-01 Drawing on their extensive research Nichols and Berliner document and categorize the ways that high stakes testing threatens the purposes and ideals of the American education system For more than a decade the debate over high stakes testing has dominated the field of education This passionate and provocative book provides a fresh perspective on the issue and powerful ammunition for opponents of high stakes tests Their analysis is grounded in the application of Campbell's Law which posits that the greater the social consequences associated with a quantitative indicator such as test scores the more likely it is that the indicator itself will become corrupted and the more likely it is that the use of the indicator will corrupt the social processes it was intended to monitor Nichols and Berliner illustrate both aspects of this corruption showing how the pressures of high stakes testing erode the validity of test scores and distort the integrity of the education system Their analysis provides a coherent and comprehensive intellectual framework for the wide ranging arguments against high stakes testing while putting a compelling human face on the data marshalled in support of those arguments

**Science Books & Films**, 1989 Practice the ASVAB! Armed Services Vocational Aptitude Battery practice test questions Complete Test Preparation Inc., 2012-10-01 Over 400 ASVAB practice questions prepared by a dedicated team of exam experts with detailed answer key exam tips and multiple choice strategies ASVAB Practice tests will help you Learn faster Practice with 2 complete ASVAB practice question sets over 400 questions Increase your score with multiple choice strategies from exam experts Learn what you MUST do in the exam room Avoid common mistakes on a test Answer multiple choice questions strategically Practice Questions include General Science Arithmetic Reasoning Word Knowledge Paragraph Comprehension Auto and Shop Mathematics Knowledge Mechanical Comprehension Electronics ASVAB Practice tests are a critical self assessment tool that reveals your strengths and weaknesses familiarize you with the exam format and types of questions build your self confidence and practice your exam time management All of these can make a huge difference in your score Practice Tests also reduce Test Anxiety one of the main reasons for low marks on an exam How to Take the ASVAB Test The Complete Guide Let's face it test taking is really not easy While some people seem to have the natural ability to know what to study how to absorb and retain information and how to stay calm enough while actually taking a test to earn a great score most of us find taking tests to be sheer misery This is one of the most important chapters Here you will find out How to Take the ASVAB Test The basics In the Test Room What you MUST do The Ultimate Guide to Test Preparation Everything you need to know Common Mistakes on a Test And how to avoid them Mental Prep How to psych yourself up for a test Why not do everything you can to get the best score on the ASVAB

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researcher enrolled as a public high school student Over the course of seventy nine consecutive days John D Butler participates in and observes Rhode Island 2013 Teacher of the Year Jessica M Waters s high school chemistry class documenting his experiences as they unfold Who s the New Kid in Chemistry is a compelling example of what can be accomplished when an educational researcher and teacher collaborate in the classroom This work includes a discussion on flexible homework assignments data driven instruction and thirty teacher best practices This book is an invaluable resource for teachers across all content areas masters and doctoral research method classes and future Teachers of the Year

**Graduate Education in the Chemical Sciences** National Research Council, Commission on Physical Sciences, Mathematics, and Applications, Board on Chemical Sciences and Technology, Chemical Sciences Roundtable, 2000-07-31 Graduate Education in the Chemical Sciences is a summary of the December 1999 workshop Graduate Education in the Chemical Sciences Issues for the 21st Century This workshop discussed the various features of graduate education in chemical science and technology Using case histories and their individual experiences speakers examined the current status of graduate education in the chemical sciences identified problems and opportunities and discussed possible strategies for improving the system The discussion was oriented toward the goal of generating graduates who are well prepared to advance the chemical sciences in academia government and industry in the next 5 to 10 years **The Science Teacher** , 1987

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