

100) A student finds the ΔG of hydrolysis of ATP is -30.5 kJ/mol at pH 7.0. Is this value the same as the ΔG of phosphorylation? At which condition would phosphorylation be most favorable?

101) What does the term "coupled reaction" mean?

_____ ΔG of 1 reaction is used to drive ΔG of 2nd reaction _____ 200

102) What is the standard free energy change of the hydrolysis of ATP?

_____ -30.5 kJ/mol _____ 200

103) Calculate the ΔG of phosphorylation of ADP to ATP at pH 7.0.

$\frac{1.2 \times 10^{-3}}{2 \times 10^{-3}} = \frac{e^{-\Delta G/RT}}{e^{-\Delta G/RT}}$ $\Delta G = -RT \ln \left(\frac{[ADP][P_i]}{[ATP]} \right)$
 $\Delta G = - (8.314 \text{ J/mol}\cdot\text{K}) (298 \text{ K}) \ln \left(\frac{1.2 \times 10^{-3}}{2 \times 10^{-3}} \right)$
 $\Delta G = - (8.314 \text{ J/mol}\cdot\text{K}) (298 \text{ K}) \ln (0.6)$
 $\Delta G = - (8.314 \text{ J/mol}\cdot\text{K}) (298 \text{ K}) (-0.5108)$
 $\Delta G = 1.26 \text{ kJ/mol}$ _____ 200

104) What is the standard free energy change of the hydrolysis of ATP at pH 7.0?

105) What is the standard free energy change of the hydrolysis of ATP at pH 7.0?

_____ -30.5 kJ/mol _____ 200

106) What is the standard free energy change of the hydrolysis of ATP at pH 7.0?

_____ -30.5 kJ/mol _____ 200

_____ -30.5 kJ/mol _____ 200

_____ -30.5 kJ/mol _____ 200

(Total: 200)



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