

Question			Answer	Marks	Guidance
5	(a)	(i)	<p><i>Nuclear charge mark</i> (Across the period) number of protons increases <b>OR</b> greater nuclear charge ✓</p> <p><i>Quality of written communication – nuclear OR proton(s) OR nucleus spelled correctly ONCE for the first marking point</i></p> <p><i>Distance / shielding mark</i> (Outermost) electrons are in the same shell <b>OR</b> (Outermost) electrons experience the same shielding <b>OR</b> Atomic radius decreases ✓</p> <p><i>Nuclear attraction (to electron) mark</i> Greater nuclear attraction (on outermost electrons) <b>OR</b> (outer) electrons are attracted more strongly (to the nucleus) ✓</p>	3	<p><b>FULL ANNOTATIONS WITH TICKS, CROSSES, CON, etc MUST BE USED</b></p> <p>Comparison should be used for each mark</p> <p><b>IGNORE</b> atomic number increases, but <b>ALLOW</b> proton number increases <b>IGNORE</b> nucleus gets bigger <b>IGNORE</b> 'effective nuclear charge increases' <b>DO NOT ALLOW</b> 'charge' increases without reference to nuclear</p> <p><b>ALLOW</b> shielding is similar BUT <b>IGNORE</b> 'there is shielding' <b>DO NOT ALLOW</b> sub-shells <b>OR</b> orbitals</p> <p><b>ALLOW</b> greater nuclear pull for greater nuclear attraction <b>DO NOT ALLOW</b> use of greater nuclear charge for greater nuclear attraction for third mark</p>
		(ii)	(Diamond and graphite form) <b>gaseous atoms</b> (of carbon when they are ionised) ✓	1	<b>ALLOW</b> the atoms are in the gaseous state

# F321 Jun 2013 Paper

**Peter Cann, Peter Hughes**



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