#### CHEMISTRY MODULE 2 – FOUNDATIONS IN CHEMISTRY

# 2.1 – ATOMS AND REACTIONS

#### Atomic structure and isotopes

 Describe the structure of an atom including the distribution of mass and charge.

	Protons	Electrons	Neutrons
Mass	1	1/2000th	1
Charge	+1	-1	0

- Mass is in nucleus(protons and neutons) with electron cloud surrounding it.
- Describe the difference between mass (nucleon) number and atomic (proton) number
  - Mass number the total number of protons and neutrons in the nucleus of an isotope
  - Atomic number the total number of protons in the nucleus of an atom
- Explain what an isotope is.
  - Atoms of an element (same number of protons) with different numbers of neutrons and different masses.
- Define the term 'first ionisation energy'
  - The amount of energy needed to remove an electron from each atom in 1 mole of gaseous atoms of an element.

#### Relative masses

- State what is used as the standard measurement for relative masses
  - o Carbon-12
- Define 'relative atomic mass' (Ar)
  - o The weighted average mass of an atom, relative to 1/12th the mass of a carbon-12 atom. THIS MEANS:

(The weighted average atomic mass of an element, relative to carbon 12, on a scale where carbon 12 is 12.)

- Define 'relative isotopic mass'
  - The mass of an isotope compared to 1/12th the mass of a carbon-12 atom.

THIS MEANS

(The mass of an isotope of an element relative to carbon 12 on a scale where carbon 12 is 12.)

# F321 Chemistry Jun 2013

**Zhe Xu** 

# F321 Chemistry Jun 2013:

The book delves into F321 Chemistry Jun 2013. F321 Chemistry Jun 2013 is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into F321 Chemistry Jun 2013, encompassing both the fundamentals and more intricate discussions.

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  - Chapter 1: Introduction to F321 Chemistry Jun 2013
  - o Chapter 2: Essential Elements of F321 Chemistry Jun 2013
  - o Chapter 3: F321 Chemistry Jun 2013 in Everyday Life
  - Chapter 4: F321 Chemistry Jun 2013 in Specific Contexts
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