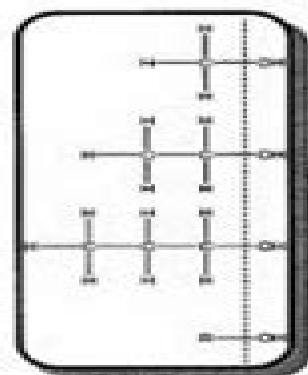


# Naming Substituted Hydrocarbons

A substituted hydrocarbon is a hydrocarbon with an element other than hydrogen attached somewhere along the hydrocarbon chain. It is named in a similar fashion to a hydrocarbon. This can be illustrated with alcohols as an example. The compounds pictured to the lower left are



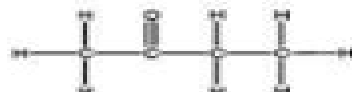
alcohols. They look like alkanes with  $-OH$  at one end where a hydrogen would have been. The  $-OH$  is called a functional group. The rest of the molecule is called a residue ( $R$ ). The general formula for alcohols is  $R-OH$ .  $CH_3OH$ , the first alcohol pictured to the left is formed by substituting an  $-OH$  group for hydrogen on methane ( $CH_4$ ). As a result, it is called 1-methanol. The suffix *ol* shows that it is an alcohol. The root *methan* comes from methane. The number 1 shows the location of the  $-OH$ . The next alcohol in the series,  $CH_3CH_2OH$ , formed from ethane, is called 1-ethanol.

$CH_3CH_2CH_2OH$  is 1-propanol.

The alcohols and several other classes of substituted hydrocarbons are found in *Table R*. The root is determined by counting the number of carbons in the chain. For halides, the substitution is identified with a prefix. For the remaining substitutions, a suffix is used. (See *Table R*.) As with all hydrocarbons, the number and location of groups needs to be identified.



1,2,2-trifluoropropane



2-butanone

Name the following compounds using the rules for naming hydrocarbons and by referring to the reading and *Table R* above.

1.  $CH_3CH_2CHOHCH_3$  \_\_\_\_\_



**Table R**  
**Organic Functional Groups**

| Class of Compound   | Functional Group                                                    | General Formula                                    | Example                                                                  |
|---------------------|---------------------------------------------------------------------|----------------------------------------------------|--------------------------------------------------------------------------|
| halide (halocarbon) | $-F$ (fluoro-)<br>$-Cl$ (chloro-)<br>$-Br$ (bromo-)<br>$-I$ (iodo-) | $R-X$<br>( $X$ represents any halogen)             | $CH_3CHClCH_3$<br>2-chloropropane                                        |
| alcohol             | $-OH$                                                               | $R-OH$                                             | $CH_3CH_2CH_2OH$<br>1-propanol                                           |
| ether               | $-O-$                                                               | $R-O-R'$                                           | $CH_3OCH_2CH_3$<br>methyl ethyl ether                                    |
| aldehyde            | $\begin{array}{c} O \\    \\ -C-H \end{array}$                      | $R-\begin{array}{c} O \\    \\ C-H \end{array}$    | $CH_3CH_2\begin{array}{c} O \\    \\ C-H \end{array}$<br>propanal        |
| ketone              | $\begin{array}{c} O \\    \\ -C- \end{array}$                       | $R-\begin{array}{c} O \\    \\ C-R' \end{array}$   | $CH_3COCH_2CH_2CH_3$<br>2-pentanone                                      |
| organic acid        | $\begin{array}{c} O \\    \\ -C-OH \end{array}$                     | $R-\begin{array}{c} O \\    \\ C-OH \end{array}$   | $CH_3CH_2\begin{array}{c} O \\    \\ C-OH \end{array}$<br>propanoic acid |
| ester               | $\begin{array}{c} O \\    \\ -C-O- \end{array}$                     | $R-\begin{array}{c} O \\    \\ C-O-R' \end{array}$ | $CH_3CH_2COCH_3$<br>methyl propanoate                                    |
| amine               | $-N-$                                                               | $R'-N-R''$                                         | $CH_3CH_2CH_2NH_2$<br>1-propanamine                                      |
| amide               | $\begin{array}{c} O \\    \\ -C-NH \end{array}$                     | $R-\begin{array}{c} O \\    \\ C-NH \end{array}$   | $CH_3CH_2\begin{array}{c} O \\    \\ C-NH_2 \end{array}$<br>propanamide  |

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