

Chemistry EOC Review – Answer Key

- 1) a. 68,200,000 cg b. 5000 mm c. 0.0125 kL d. 0.548 dm
- 2) ~21.5 mL
- 3) production of heat & light, formation of a gas, formation of a precipitate, color change
- 4) $33.64 \text{ g} / 4.28 \text{ ml} = 7.86 \text{ g/ml} = \text{Fe}$
- 5) m.p. $-183^{\circ}\text{C} = \text{CH}_4$ (methane)
- 6) H_2 – splint test (burning splint ignites hydrogen)
 O_2 – splint test (oxygen reignites glowing splint)
 H_2O – cobalt chloride paper (paper changes from blue to pink in the presence of water vapor)
 CO_2 – splint test (carbon dioxide extinguishes burning splint);
limewater test (limewater turns cloudy when carbon dioxide is bubbled through it,
producing CaCO_3)
- 7) The protons are positively charged particles found inside the nucleus along with the neutrons, which have no charge. The total number of these nucleons in an atom is the mass # / atomic mass of the atom. The electrons are the negatively charged particles found orbiting in “clouds” around the nucleus. Electrons can be gained by the atom to form anions or lost to form cations.
- 8) The thing that makes any element on the periodic table different from all other elements is the number of protons in its nucleus. Isotopes are atoms of the same type of element which have different masses because they have different numbers of neutrons.
- 9) a. $^{39}_{19}\text{K}^+$
b. $^{57}_{26}\text{Fe}$
c. $^{33}_{16}\text{S}^{2-}$
- 10)
- | <u>chemical symbol</u> | <u>atomic number</u> | <u>mass number</u> | <u># of protons</u> | <u># of electrons</u> | <u># of neutrons</u> | <u>net charge</u> |
|------------------------|----------------------|--------------------|---------------------|-----------------------|----------------------|-------------------|
| K^+ | 19 | 42 | 19 | 18 | 23 | +1 |
| Sr | 38 | 88 | 38 | 38 | 50 | 0 |
| F^- | 9 | 20 | 9 | 10 | 11 | -1 |
| P | 15 | 31 | 15 | 15 | 16 | 0 |
| Ca^{2+} | 20 | 42 | 20 | 18 | 22 | +2 |
- 11) uranium – 238 $238 - 92 = 146$ neutrons
- 12) (1) 199.59g
(2) They contain the same # of protons (they are the same element), but different #s of neutrons (electrons don't matter, b/c they have virtually no mass), & thus different masses.

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Table of Contents Eoc Chemistry Practice Key

1. Understanding the eBook Eoc Chemistry Practice Key
 - The Rise of Digital Reading Eoc Chemistry Practice Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Eoc Chemistry Practice Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Eoc Chemistry Practice Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Eoc Chemistry Practice Key
 - Personalized Recommendations
 - Eoc Chemistry Practice Key User Reviews and Ratings
 - Eoc Chemistry Practice Key and Bestseller Lists
5. Accessing Eoc Chemistry Practice Key Free and Paid eBooks

- Eoc Chemistry Practice Key Public Domain eBooks
 - Eoc Chemistry Practice Key eBook Subscription Services
 - Eoc Chemistry Practice Key Budget-Friendly Options
6. Navigating Eoc Chemistry Practice Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Eoc Chemistry Practice Key Compatibility with Devices
 - Eoc Chemistry Practice Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Eoc Chemistry Practice Key
 - Highlighting and Note-Taking Eoc Chemistry Practice Key
 - Interactive Elements Eoc Chemistry Practice Key
 8. Staying Engaged with Eoc Chemistry Practice Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Eoc Chemistry Practice Key
 9. Balancing eBooks and Physical Books Eoc Chemistry Practice Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Eoc Chemistry Practice Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Eoc Chemistry Practice Key
 - Setting Reading Goals Eoc Chemistry Practice Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Eoc Chemistry Practice Key
 - Fact-Checking eBook Content of Eoc Chemistry Practice Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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