

Quiz 4 – Potential Energy Diagrams

1. The addition of a catalyst to a reaction provides an alternative mechanism with

- A. Lower activation energy and lower reaction rate
- B. Lower activation energy and higher reaction rate**
- C. Higher activation energy and lower reaction rate
- D. Higher activation energy and higher reaction rate

2. Consider the following Reaction



The correct equation including the heat term is

- A. $\text{N}_{2(g)} + \text{O}_{2(g)} + 90 \text{ kJ} \longrightarrow 2\text{NO}_{(g)}$
- B. $\text{N}_{2(g)} + \text{O}_{2(g)} + 180 \text{ kJ} \longrightarrow 2\text{NO}_{(g)}$**
- C. $\text{N}_{2(g)} + \text{O}_{2(g)} \longrightarrow 2\text{NO}_{(g)} + 90 \text{ kJ}$
- D. $\text{N}_{2(g)} + \text{O}_{2(g)} \longrightarrow 2\text{NO}_{(g)} + 180 \text{ kJ}$

3. A forward reaction has an activation energy of 50 kJ and a ΔH of -100 kJ .
The PE diagram, which describes this reaction, is

A.

4. Consider the following potential energy diagram

The Activation energy for the forward reaction is

- A. 25 kJ
- B. 50 kJ
- C. 75 kJ**
- D. 125 kJ

5. Consider the following reaction



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