

PHYSICAL CHEMISTRY

(Complete Syllabus)

SECTION - A

Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

- The mass of oxygen required to produce enough CO which completely reduces x gm of Fe_2O_3 . Suppose molecular weight of $\text{O}_2 = y$, molecular weight of $\text{Fe}_2\text{O}_3 = z$
(1) $\frac{2xz}{3y}$ (2) $\frac{3xy}{2z}$
(3) $\frac{2y}{3xz}$ (4) $\frac{3z}{2xy}$
- The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be
(1) 0.02 N_0 (2) 0.16 N_0
(3) 0.32 N_0 (4) 0.08 N_0
- A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with Na_2CO_3 to precipitate the calcium as calcium carbonate. This CaCO_3 is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of CaCl_2 in the original mixture is
(1) 15.2%
(2) 32.1%
(3) 21.8%
(4) 11.07%
- 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of I_2 liberated is
(1) 25.4 g (2) 12.7 g
(3) 2.54 g (4) 1.27 g
- x_1 ml solution having normality 1 is diluted to get a solution of normality y_2 . What would be volume required for dilution?
(1) $x_1 \left(1 - \frac{y_1}{y_2}\right)$ (2) $x_1 \left(\frac{y_2}{y_1} - 1\right)$
(3) $x_1 \left(\frac{y_1}{y_2} - 1\right)$ (4) $x_1 \left(1 + \frac{y_1}{y_2}\right)$
- 1.625 g of an interhalogen compound ICl_x was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is
(1) ICl (2) ICl_3
(3) ICl_5 (4) ICl_7
- If 12.25 g of a sample of KClO_3 of 75% purity and rest vapourisable impurities is strongly heated to give KCl , KClO_4 and 0.32 gm O_2 . The mass of residue obtained is
(1) 1.769 g (2) 7.096 g
(3) 5.327 g (4) 8.85 g
- A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is
(1) XeF_2 (2) XeF_4
(3) XeF_6 (4) XeF_8
- 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl . The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

Multiple Choice Physical Chemistry Exam Questions

Zhe Xu



Multiple Choice Physical Chemistry Exam Questions:

The book delves into Multiple Choice Physical Chemistry Exam Questions. Multiple Choice Physical Chemistry Exam Questions is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Multiple Choice Physical Chemistry Exam Questions, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Multiple Choice Physical Chemistry Exam Questions
 - Chapter 2: Essential Elements of Multiple Choice Physical Chemistry Exam Questions
 - Chapter 3: Multiple Choice Physical Chemistry Exam Questions in Everyday Life
 - Chapter 4: Multiple Choice Physical Chemistry Exam Questions in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, the author will provide an overview of Multiple Choice Physical Chemistry Exam Questions. The first chapter will explore what Multiple Choice Physical Chemistry Exam Questions is, why Multiple Choice Physical Chemistry Exam Questions is vital, and how to effectively learn about Multiple Choice Physical Chemistry Exam Questions.
 3. In chapter 2, the author will delve into the foundational concepts of Multiple Choice Physical Chemistry Exam Questions. This chapter will elucidate the essential principles that need to be understood to grasp Multiple Choice Physical Chemistry Exam Questions in its entirety.
 4. In chapter 3, the author will examine the practical applications of Multiple Choice Physical Chemistry Exam Questions in daily life. The third chapter will showcase real-world examples of how Multiple Choice Physical Chemistry Exam Questions can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Multiple Choice Physical Chemistry Exam Questions in specific contexts. This chapter will explore how Multiple Choice Physical Chemistry Exam Questions is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Multiple Choice Physical Chemistry Exam Questions. The final chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Multiple Choice Physical Chemistry Exam Questions.

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