

Question 4

- 4(a)(i) Two ions present in the electrolyte are $\text{Na}^+_{(\text{aq})}$ (sodium) and $\text{Cl}^-_{(\text{aq})}$ (chloride).
- 4(a)(ii) Two factors that affect the discharge of an ion are position in the electrochemical series and concentration.
- 4(a)(iii) $\text{Cl}^-_{(\text{aq})}$ and $\text{OH}^-_{(\text{aq})}$ will drift towards the anode.
- 4(a)(iv) Anode reaction: $2\text{Cl}^-_{(\text{aq})} \rightarrow \text{Cl}_{2(\text{g})} + 2\text{e}^-$.
- 4(a)(v) Cathode reaction: $2\text{H}^+_{(\text{aq})} + 2\text{e}^- \rightarrow \text{H}_{2(\text{g})}$.
- 4(b) Magnesium ribbon: conductor.
Plastic ruler: non-conductor.
 $\text{MgCl}_{2(\text{aq})}$: conductor.
- 4(c) Carbon is used as a lubricant, in the form of graphite and for the tips of drill bits for drilling of oil wells, in the form of diamond.

Question 5

- 5(a)(i) Name of compound C: methyl propanoate.

Homologous series: the esters.



- 5(b)(i) The name of compound D is ethanoic acid. It is produced by heating ethanol under reflux with an excess of $\text{H}^+_{(\text{aq})}/\text{K}_2\text{Cr}_2\text{O}_{7(\text{aq})}$.
- 5(b)(ii) Fermentation is the conversion of carbohydrates to ethanol and carbon dioxide by enzymes present in yeast. Raw materials are yeast and a carbohydrate source such as fruit juice. The equation for this process is $\text{C}_6\text{H}_{12}\text{O}_{6(\text{aq})} \rightarrow 2\text{C}_2\text{H}_5\text{OH}_{(\text{aq})} + 2\text{CO}_{2(\text{g})}$. The reaction is carried out under anaerobic conditions at temperatures between 25 °C and 35 °C and at a pH of between 6 and 7.
- 5(b)(iii) One other use of ethanol is as a disinfectant/sterilizing agent.

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