

## A-LEVEL CHEMISTRY

## MARK SCHEME

## Practice Paper 1

Maximum marks: 105

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1	(a)	(i)	$[H^+][OH^-]$ <b>OR</b> $[H_3O^+][OH^-]$ <i>Ignore (aq)</i> <i>Must have [ ] not ( )</i>	1
		(ii)	$\sqrt{3.46 \times 10^{-14}}$ ( $= 1.86 \times 10^{-7}$ ) <i>If no square root, CE= 0</i> <b>pH = 6.73</b> <b>Must be 2dp</b>	2
		(iii)	$[H^+] = 10^{-11.26}$ ( $= 4.365 \times 10^{-12}$ <b>OR</b> $4.37 \times 10^{-12}$ ) <i>Mark for working</i> $K_w = [4.365 \times 10^{-12} \text{ OR } 4.37 \times 10^{-12} \times 0.047] = 2.05 \times 10^{-13}$ <i>Allow <math>2.05 \times 10^{-13} = 2.1 \times 10^{-13}</math></i> <i>Mark for answer</i> <i>Ignore units</i>	1
	(b)	(i)	$HCOOH \rightleftharpoons HCOO^- + H^+$ <i>Must have <math>\rightleftharpoons</math> but ignore brackets.</i> <b>OR</b> $HCOOH + H_2O \rightleftharpoons HCOO^- + H_3O^+$ <i>Allow <math>HCO_2^-</math> or <math>CHOO^-</math> ie minus must be on oxygen, so penalise <math>COOH^-</math></i>	1
		(ii)	$K_a = \frac{[H^+][HCOO^-]}{[HCOOH]}$ <b>OR</b> $\frac{[H_3O^+][HCOO^-]}{[HCOOH]}$ <i>Must have all brackets but allow ( )</i> <i>Must be <math>HCOOH</math> etc.</i> <i>Allow ecf in formulae from (b)(i)</i>	1
		(iii)	<b>M1</b> $K_a = \frac{[H^+]^2}{[HCOOH]}$ ( $[H^+]^2 = 1.78 \times 10^{-4} \times 0.056 = 9.97 \times 10^{-6}$ ) <i>Allow HA or HX etc.</i> <i>Allow <math>[H^+] = \sqrt{K_a \times [HA]}</math> for M1</i>	1
			<b>M2</b> $[H^+] = 3.16 \times 10^{-3}$ <i>Mark for answer</i>	1
			<b>M3</b> <b>pH = 2.50</b> allow more than 2 dp but not fewer <i>Allow correct pH from their wrong <math>[H^+]</math> here only if square root shown but not taken, pH = 5.00 can score max 2 for M1 and M3</i>	1

# May 200a Level Chemistry Paper

**David Baud**



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