

Sulfur clock lab

Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1 mL	0.1 M H_2O_2 , 0.01 M KSCN	Reaction setup
B	0.5 mL	0.1 M H_2O_2 , 0.01 M KSCN	0.5 mL, 0.01 M KSCN
C	0.25 mL	0.1 M H_2O_2 , 0.01 M KSCN	0.25 mL, 0.01 M KSCN
D	0.125 mL	0.1 M H_2O_2 , 0.01 M KSCN	0.125 mL, 0.01 M KSCN

4. Add the solution using a graduated cylinder with completely standard chemical balance the reaction solution.
5. Transfer exactly the solution into a 100 mL beaker in 10 mL using a graduated cylinder. Mixing for 10 min with 10 mL of distilled water. Transfer 10 mL to 100 mL.
6. Add exactly 10 mL of distilled water. Transfer 10 mL to 100 mL.
7. Prepare and use the standard solution.
8. Prepare solution B, C, D, E as shown:
 - a. Transfer 0.5 mL, 0.25 mL, 0.125 mL, 0.0625 mL of solution A to beaker
 - b. Add 10 mL of 0.01 M KSCN solution using a graduated cylinder
 - c. Add 10 mL of 0.1 M H_2O_2 solution using a graduated cylinder
 - d. Add 10 mL of 0.01 M KSCN solution using a graduated cylinder
9. Transfer solution A into the graduated cylinder in 100 mL beaker. Use the same method with any other solution.

Lab Sulfur Clock Answers

Judd E. Hollander



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