



THE KENYA NATIONAL EXAMINATIONS COUNCIL

KNEC/GEN/TD/PSE/BUS/TT/30/2019

2019 JULY BUSINESS EXAMINATIONS TIMETABLE

BUSINESS EDUCATION SINGLE AND GROUP CERTIFICATE (S & G SUBJECTS) BUSINESS TECHNICAL EDUCATION PROGRAMME (BTEP) TECHNICAL AND VOCATIONAL EDUCATION AND TRAINING (TVET) BUSINESS MANAGEMENT COURSES (BMC)

1.0 GENERAL INSTRUCTIONS

- 1.1 The time allowed for each paper is indicated against the name of the paper and **NO EXTRA TIME IS TO BE ALLOWED**. In case of any discrepancy, the time stated on the question paper should be taken as correct.
- 1.2 Visually handicapped candidates will NOT be given extra time other than what is stated on their respective papers. They are allowed 30 minutes over and above the times indicated on this timetable for all subjects.
- 1.3 Any time allocated to reading through questions is included in the total time shown, except when special paper instructions indicate otherwise.
- 1.4 **IMPORTANT NOTICE:**
 - 1.4.1 Your attention is drawn to the regulations governing the conduct of candidates on page 20.
 - 1.4.2 Supervisors **MUST** ensure that candidates have written their names and index numbers on their answer scripts before they take the scripts from candidates.

2.0 EXAMINATION PERIOD AND DATES

2.1 ORAL PAPERS/ PROJECTS

DAY & DATE	EXAMINATION	SUBJECT CODE & NAME			DURATION
WEDNESDAY 10 TH TO FRIDAY 12 TH JULY 2019	S & C	1.	1056	Foundations of French (Oral)	30 min
		2.	1918/103/2, 1919/103/2	French I (Oral)	45 min
		3.	1918/106/2, 1919/106/2	German I (Oral)	30 min
		4.	1918/204/2, 1919/204/2	French II (Oral)	45 min
		5.	1918/208/2, 1919/208/2	German II (Oral)	30 min
		6.	2812/304/2	German Paper 2 (Oral)	30 min
		7.	2812/308/2	French Paper 2 (Oral)	30 min
		8.	2912/107/2, 2918/107/2	French I (Oral)	30 min
		9.	2912/109/2, 2918/109/2	German I (Oral)	30 min
		10.	2912/206/2, 2918/206/2	French II (Oral)	45 min
		11.	2912/208/2, 2918/208/2	German II (Oral)	30 min
		12.	2912/305/2, 2918/305/2	French III (Oral)	30 min
		13.	2912/307/2, 2918/307/2	German III (Oral)	30 min
Dates of submission will be communicated to the Institutions by e-mail.	BTEP/TVET	14.	207, 306 & 307	BTEP Craft, Diploma & Higher Diploma	Upload of Projects & Course Work Assessment Marks
		15.	107 & 207	TVET Craft	
		16.	107/108 & 307/308	TVET Diploma	
FRIDAY 12 TH JULY 2019	ALL	17.	REHEARSAL		

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Julia Schneider



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the first step to writing a net ionic equation is balancing the chemical equation present let's use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^- aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{ Ag}^+ \text{aq} + 2 \text{ F}^- \text{aq} + 2 \text{ NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2 \text{ NH}_4^+ \text{aq} + 2 \text{ F}^- \text{aq}$ because two NH_4^+ aq

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