

Redox Exercise (Part 2)

Chapter 20 Worksheet: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.



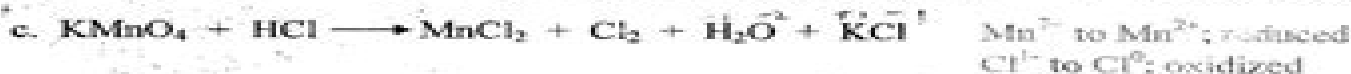
11. Give the oxidation number of each kind of atom or ion.

a. sulfate	b. Sn	c. S^{2-}	d. Fe^{3+}	e. Sn^{4+}	f. nitrate	g. ammonium
2-	0	2-	3+	4+	1-	1+

12. Calculate the oxidation number of chromium in each of the following.

a. Cr_2O_3	b. $\text{Na}_2\text{Cr}_2\text{O}_7$	c. CrSO_4	d. chromate	e. dichromate
3+	6+	2+	6+	6+

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)



14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- | | |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur | reducing agent: carbon |
| b. oxidizing agent: nitrogen | reducing agent: iodine |
| c. oxidizing agent: manganese | reducing agent: chloride |
| d. oxidizing agent: nitrogen | reducing agent: antimony |

Homework Oxidation And Reduction Answer Key

Shasha Hu



Homework Oxidation And Reduction Answer Key:

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Table of Contents Homework Oxidation And Reduction Answer Key

1. Understanding the eBook Homework Oxidation And Reduction Answer Key
 - The Rise of Digital Reading Homework Oxidation And Reduction Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Homework Oxidation And Reduction Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Homework Oxidation And Reduction Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Homework Oxidation And Reduction Answer Key
 - Personalized Recommendations
 - Homework Oxidation And Reduction Answer Key User Reviews and Ratings
 - Homework Oxidation And Reduction Answer Key and Bestseller Lists
5. Accessing Homework Oxidation And Reduction Answer Key Free and Paid eBooks
 - Homework Oxidation And Reduction Answer Key Public Domain eBooks
 - Homework Oxidation And Reduction Answer Key eBook Subscription Services
 - Homework Oxidation And Reduction Answer Key Budget-Friendly Options
6. Navigating Homework Oxidation And Reduction Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Homework Oxidation And Reduction Answer Key Compatibility with Devices
 - Homework Oxidation And Reduction Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Homework Oxidation And Reduction Answer Key
 - Highlighting and Note-Taking Homework Oxidation And Reduction Answer Key
 - Interactive Elements Homework Oxidation And Reduction Answer Key
 8. Staying Engaged with Homework Oxidation And Reduction Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Homework Oxidation And Reduction Answer Key
 9. Balancing eBooks and Physical Books Homework Oxidation And Reduction Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Homework Oxidation And Reduction Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Homework Oxidation And Reduction Answer Key
 - Setting Reading Goals Homework Oxidation And Reduction Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Homework Oxidation And Reduction Answer Key
 - Fact-Checking eBook Content of Homework Oxidation And Reduction Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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