

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	2. O_2	2. 4 molecules NO
1.00 mol	2.00 mol	2. 2.00 mol	3. none	3. none
4.00 mol	6.00 mol	4. 4.00 mol	4. NO	4. 2.00 mol O_2
6.00 mol	1.00 mol	6. 1.00 mol	6. NO	6. 1.50 mol O_2
1.00 mol	6.00 mol	8. 2.00 mol	8. O_2	8. 2.00 mol NO
6.000 mol	6.200 mol	10. 6.200 mol	10. NO	10. 0.400 mol O_2
32.00 g	60.02 g	12. 92.02 g	12. none	12. none
64.00 g	60.02 g	14. 66.01 g	14. O_2	14. 30.12 g NO
64.00 g	30.01 g	16. 28.76 g	16. O_2	17. 1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 1.302 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of $\text{NO}:\text{O}_2$ is larger than the balanced equation mole ratio of $\text{NO}:\text{O}_2$, there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

Excess $\text{NO} = 20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

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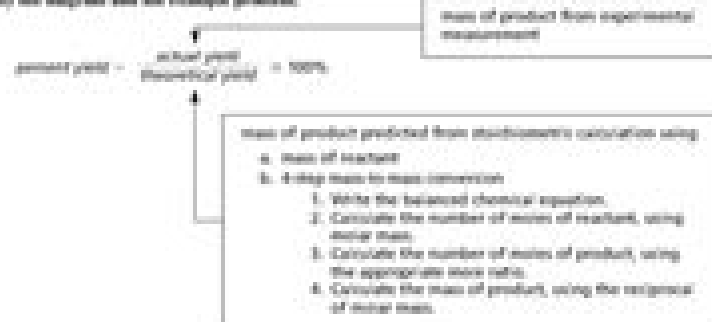
CHAPTER 13

STUDY GUIDE FOR CONTENT MASTERY

Section 13.4 Percent Yield

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 1 \text{ mol } \text{Ga}$ 7.00 g Ga $\times$ 1 mol Ga/69.72 g Ga
- number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}} = 0.200 \text{ mol } \text{Ga}_2\text{O}_3$ 0.100 mol Ga $\times$ 2 mol Ga_2O_3 /1 mol Ga
- theoretical yield: $0.200 \text{ mol } \text{Ga}_2\text{O}_3 \times \frac{157.44 \text{ g } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}_2\text{O}_3} = 31.49 \text{ g } \text{Ga}_2\text{O}_3$ correct
- percent yield: $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{31.49 \text{ g } \text{Ga}_2\text{O}_3} \times 100 = 23.56\%$ 7.42 g Ga_2O_3 /31.49 g Ga_2O_3 $\times$ 100

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Table of Contents Glencoe Mcgraw Hill Chemistry Answers

1. Understanding the eBook Glencoe Mcgraw Hill Chemistry Answers
 - The Rise of Digital Reading Glencoe Mcgraw Hill Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Mcgraw Hill Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Glencoe Mcgraw Hill Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Mcgraw Hill Chemistry Answers
 - Personalized Recommendations
 - Glencoe Mcgraw Hill Chemistry Answers User Reviews and Ratings
 - Glencoe Mcgraw Hill Chemistry Answers and Bestseller Lists
5. Accessing Glencoe Mcgraw Hill Chemistry Answers Free and Paid eBooks
 - Glencoe Mcgraw Hill Chemistry Answers Public Domain eBooks
 - Glencoe Mcgraw Hill Chemistry Answers eBook Subscription Services
 - Glencoe Mcgraw Hill Chemistry Answers Budget-Friendly Options
6. Navigating Glencoe Mcgraw Hill Chemistry Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Glencoe Mcgraw Hill Chemistry Answers Compatibility with Devices
 - Glencoe Mcgraw Hill Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Glencoe Mcgraw Hill Chemistry Answers
 - Highlighting and Note-Taking Glencoe Mcgraw Hill Chemistry Answers
 - Interactive Elements Glencoe Mcgraw Hill Chemistry Answers
8. Staying Engaged with Glencoe Mcgraw Hill Chemistry Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glencoe Mcgraw Hill Chemistry Answers
9. Balancing eBooks and Physical Books Glencoe Mcgraw Hill Chemistry Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glencoe Mcgraw Hill Chemistry Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Glencoe Mcgraw Hill Chemistry Answers
- Setting Reading Goals Glencoe Mcgraw Hill Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Glencoe Mcgraw Hill Chemistry Answers
- Fact-Checking eBook Content of Glencoe Mcgraw Hill Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
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