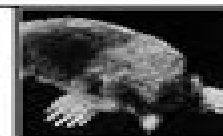


Activity A:**Moles**Get the Gizmo ready:

- Check that the equation is still
$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$
- If not, click **New equation** until it reappears.



Introduction: A **mole** is:

- A) A mammal known for digging up gardens.
- B) A small, dark spot on the skin.
- C) A spy embedded within an enemy government.
- D) 6.02×10^{23} particles of a substance.
- E) All of the above.

The correct answer, of course, is E. In chemistry, the mole (mol) is defined as an amount of a substance that contains 6.022×10^{23} particles of that substance. This number, called **Avogadro's number**, is special because this number of particles has a mass in grams that is equal to the mass (in unified mass units) of a single particle of the substance.

Question: How do scientists find the formula mass and molar mass of a substance?

1. 1. **Calculate:** The **formula mass** of a compound is the sum of the masses of each atom in the chemical formula. The unit of formula mass is the unified mass unit (u). Formula mass is also called **molecular mass** if the compound is composed of molecules.

Iron's atomic mass is 55.85 u, carbon's mass is 12.01 u, and oxygen's mass is 16.00 u.

- A. Calculate the formula mass of carbon monoxide (CO) by adding the atomic mass of carbon and the atomic mass of oxygen:

- B. Calculate the formula mass of carbon dioxide (CO₂):

- C. Calculate the formula mass of iron (III) oxide (Fe₂O₃):

2. **Inter:** A mole of a substance has a mass in grams that is equal to the formula mass. For example, a carbon atom has an average mass of 12.01 u. A mole of carbon has a mass of 12.01 g. Based on their formula masses, list the **molar mass** of each substance. The unit for molar mass is g/mol, or grams per mole.

Fe₂O₃ CO Fe CO₂

Check your answers on the Gizmo by inspecting the middle row of tiles on the right side of the Gizmo. These tiles show the units "1 mol" on top and "g" below.

3. **Practice:** Hydrogen has an atomic mass of 1.01 u. What is the molar mass of these substances? (Remember to use the units g/mol.)

H₂O CH₄ H₂CO₃ C₆H₁₂O₆

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D Keegan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'D Keegan'.

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Table of Contents Gizmo Stoichiometry Answers Activity A

1. Understanding the eBook Gizmo Stoichiometry Answers Activity A
 - The Rise of Digital Reading Gizmo Stoichiometry Answers Activity A
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmo Stoichiometry Answers Activity A
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmo Stoichiometry Answers Activity A
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmo Stoichiometry Answers Activity A
 - Personalized Recommendations
 - Gizmo Stoichiometry Answers Activity A User Reviews and Ratings
 - Gizmo Stoichiometry Answers Activity A and Bestseller Lists
5. Accessing Gizmo Stoichiometry Answers Activity A Free and Paid eBooks
 - Gizmo Stoichiometry Answers Activity A Public Domain eBooks
 - Gizmo Stoichiometry Answers Activity A eBook Subscription Services
 - Gizmo Stoichiometry Answers Activity A Budget-Friendly Options

6. Navigating Gizmo Stoichiometry Answers Activity A eBook Formats
 - ePub, PDF, MOBI, and More
 - Gizmo Stoichiometry Answers Activity A Compatibility with Devices
 - Gizmo Stoichiometry Answers Activity A Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmo Stoichiometry Answers Activity A
 - Highlighting and Note-Taking Gizmo Stoichiometry Answers Activity A
 - Interactive Elements Gizmo Stoichiometry Answers Activity A
8. Staying Engaged with Gizmo Stoichiometry Answers Activity A
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmo Stoichiometry Answers Activity A
9. Balancing eBooks and Physical Books Gizmo Stoichiometry Answers Activity A
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmo Stoichiometry Answers Activity A
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gizmo Stoichiometry Answers Activity A
 - Setting Reading Goals Gizmo Stoichiometry Answers Activity A
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gizmo Stoichiometry Answers Activity A
 - Fact-Checking eBook Content of Gizmo Stoichiometry Answers Activity A
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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