

Part B-1

Answer all questions in this part.

Directions (31–50): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

31 What is the number of electrons in an Al^{3+} ion?

- (1) 10 (3) 3
(2) 13 (4) 16

Al +13p

32 The valence electron of which atom in the ground state has the greatest amount of energy?

- (1) cesium (3) rubidium
(2) lithium (4) sodium

33 The numbers of protons and neutrons in each of four different atoms are shown in the table below.

Protons and Neutrons in Four Different Atoms

Atom	Number of Protons	Number of Neutrons
A	8	8
D	9	9
E	9	10
G	10	10

Which two atoms represent isotopes of the same element?

- (1) A and D (3) E and D
(2) A and G (4) E and G

34 Which elements have the most similar chemical properties?

- (1) boron and carbon
(2) oxygen and sulfur
(3) aluminum and bromine
(4) argon and silicon

35 Which element reacts with oxygen to form ionic bonds?

- (1) calcium (3) chlorine
(2) hydrogen (4) nitrogen

36 The table below gives the atomic mass and the abundance of the two naturally occurring isotopes of chlorine.

Naturally Occurring Isotopes of Chlorine

Isotopes	Atomic Mass of the Isotope (u)	Natural Abundance (%)
^{35}Cl	34.97	75.76
^{37}Cl	36.97	24.24

Which numerical setup can be used to calculate the atomic mass of the element chlorine?

- (1) $(34.97 \text{ u})(75.76) + (36.97 \text{ u})(24.24)$
(2) $(34.97 \text{ u})(0.2424) + (36.97 \text{ u})(0.7576)$
(3) $(34.97 \text{ u})(0.7576) + (36.97 \text{ u})(0.2424)$
(4) $(34.97 \text{ u})(24.24) + (36.97 \text{ u})(75.76)$

37 Which general trends in first ionization energy and electronegativity values are demonstrated by Group 15 elements as they are considered in order from top to bottom?

- (1) The first ionization energy decreases and the electronegativity decreases.
(2) The first ionization energy increases and the electronegativity increases.
(3) The first ionization energy decreases and the electronegativity increases.
(4) The first ionization energy increases and the electronegativity decreases.

38 An aluminum sample has a mass of 80.01 g and a density of 2.70 g/cm³. According to the data, to what number of significant figures should the calculated volume of the aluminum sample be expressed?

- (1) 1 (3) 3
(2) 2 (4) 4

$$D = \frac{M}{V}$$

$$V = \frac{M}{D}$$

$$\frac{80.01 \text{ (4)}}{2.70 \text{ (3)}}$$

Geometry Regents Answers For June 2014

Andre, Ph.D. Castagna



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Regents Exams and Answers: Geometry, Sixth Edition Barron's Educational Series, Andre, Ph.D. Castagna, 2025-01-07
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This book is addressed to people with research interests in the nature of mathematical thinking at any level to people with an interest in higher order thinking skills in any domain and to all mathematics teachers. The focal point of the book is a framework for the analysis of complex problem solving behavior. That framework is presented in Part One which consists of Chapters 1 through 5. It describes four qualitatively different aspects of complex intellectual activity: cognitive resources, the body of facts and procedures at one's disposal, heuristics, rules of thumb for making progress in difficult situations, control having to do with the efficiency with which individuals utilize the knowledge at their disposal, and belief systems, one's perspectives regarding the nature of a discipline and how one goes about working in it. Part Two of the book, consisting of Chapters 6 through 10, presents a series of empirical studies that flesh out the analytical framework. These studies document the ways that competent problem solvers make the most of the knowledge at their disposal. They include observations of students indicating some typical roadblocks to success. Data taken from students before and after a series of intensive problem solving courses document the kinds of learning that can result from carefully designed instruction. Finally, observations made in typical high school classrooms serve to indicate some of the sources of students' often counterproductive mathematical behavior.

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