




Triangle	Trigonometric Formulae	
	Missing side	Missing angle
<div>SOH</div> 	<p>To find the length of the opposite side:</p> $O = H \times \sin(\theta)$ <p>To find the length of the hypotenuse:</p> $H = \frac{O}{\sin(\theta)}$	<p>Make θ the subject of the sine function</p> $\sin(\theta) = \frac{O}{H}$ <p>to find the missing angle using the inverse sine function</p> $\theta = \sin^{-1}\left(\frac{O}{H}\right)$
<div>CAH</div> 	<p>To find the length of the adjacent side:</p> $A = H \times \cos(\theta)$ <p>To find the length of the hypotenuse:</p> $H = \frac{A}{\cos(\theta)}$	<p>Make θ the subject of the cosine function</p> $\cos(\theta) = \frac{A}{H}$ <p>to find the missing angle using the inverse cosine function</p> $\theta = \cos^{-1}\left(\frac{A}{H}\right)$
<div>TOA</div> 	<p>To find the length of the opposite side:</p> $O = A \times \tan(\theta)$ <p>To find the length of the adjacent side:</p> $A = \frac{O}{\tan(\theta)}$	<p>Make θ the subject of the tangent function</p> $\tan(\theta) = \frac{O}{A}$ <p>to find the missing angle using the inverse tangent function</p> $\theta = \tan^{-1}\left(\frac{O}{A}\right)$

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20 balancing chemical equations worksheets answers - Nov 30 2022

web a balancing chemical equation worksheet is a practice booklet with unsolved and solved chemical equation problems on which students can practice their balancing skills if all the problems are unsolved it will usually contain an answer section

that students can use to gauge their performance

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web balancing chemical equations worksheet answer key balance the equations below

$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3 \\ 2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2 \\ 2 \text{ NaCl} + \text{F}_2 \rightarrow 2 \text{ NaF} + \text{Cl}_2 \\ 4 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O} \\ 5 \text{ Pb} + 2 \text{ OH}^- + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2 \\ 6 \text{ Al} + 3 \text{ Br}_2 \rightarrow 3 \text{ K}_2\text{SO}_4 + 6 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 + 3 \text{ H}_2\text{O} \\ 1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O} \end{array}$$

5 1 1 practice problems writing and balancing chemical equations - May 05 2023

web may 28 2020 write a balanced chemical equation for each step of the process the first step is the decomposition of solid calcium carbonate from seashells to form solid calcium oxide and gaseous carbon dioxide the second step is the formation of solid calcium hydroxide as the only product from the reaction of the solid calcium oxide with liquid water