

Grade 12 Chemistry Exam Review

Below is a list of major topics covered on the exam for this course. If a topic is not listed, it will NOT be on the exam.

Your studying should begin with your class notes, assigned problems, textbook problems and extra review questions suggested below. This list is a guide and some specific questions on the exam may not appear here.

Good luck ☺

Unit #1 – Energy Changes Text Reference: Chapter 5

- ☐ Energy changes in reactions (system, surroundings, endothermic vs. exothermic)
- ☐ Specific Heat Capacity
- ☐ Calorimetry, $q = mc\Delta T$
- ☐ Molar enthalpies
- ☐ Representing Enthalpy Changes (Thermochemical Equations, Potential Energy Diagrams, etc.)
- ☐ Hess' Law
- ☐ Standard Enthalpies of Formation

Sample Exam Questions

1. Given the following information:
 $2\text{CH}_3\text{OH}_{(l)} + 3\text{O}_{2(g)} \rightarrow 2\text{CO}_{2(g)} + 4\text{H}_2\text{O}_{(l)} \quad \Delta H^\circ = -1275 \text{ kJ}\cdot\text{mol}^{-1}$
 $\Delta H_f^\circ \text{CO}_{2(g)} = -394 \text{ kJ}\cdot\text{mol}^{-1}$, $\Delta H_f^\circ \text{H}_2\text{O}_{(l)} = -286 \text{ kJ}\cdot\text{mol}^{-1}$
 What is the standard heat formation, ΔH_f° ($\text{kJ}\cdot\text{mol}^{-1}$) of $\text{CH}_3\text{OH}_{(l)}$?
 A. -602 B. -1329 C. 1329 D. 602
2. When 729 g of liquid acetone is cooled from 22.7 °C to 25.0 °C it loses 214 kJ of heat. What is the specific heat capacity of liquid acetone?
 A. 0.213 B. 4.19 C. 0.0047 D. 3.06 E. 8.70
3. Identify the correct expression that can be used to determine the change in standard energy for the following reaction:
 $\text{C}_2\text{H}_{5(g)} + \frac{13}{2}\text{O}_{2(g)} \rightarrow 4\text{CO}_{2(g)} + 5\text{H}_2\text{O}_{(g)} + \text{heat}$
 A. $\Delta H^\circ_r = [4 \text{ mol}](\Delta H^\circ_f \text{CO}_{2(g)}) + (5 \text{ mol})(\Delta H^\circ_f \text{H}_2\text{O}_{(g)}) - [2 \text{ mol}](\Delta H^\circ_f \text{C}_2\text{H}_{5(g)}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ_f \text{O}_{2(g)})$ B. $\Delta H^\circ_r = [2 \text{ mol}](\Delta H^\circ_f \text{C}_2\text{H}_{5(g)}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ_f \text{O}_{2(g)}) - [4 \text{ mol}](\Delta H^\circ_f \text{CO}_{2(g)}) + (5 \text{ mol})(\Delta H^\circ_f \text{H}_2\text{O}_{(g)})$
 C. $\Delta H^\circ_r = [2 \text{ mol}](\Delta H^\circ_f \text{C}_2\text{H}_{5(g)}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ_f \text{O}_{2(g)}) - [4 \text{ mol}](\Delta H^\circ_f \text{CO}_{2(g)}) + (5 \text{ mol})(\Delta H^\circ_f \text{H}_2\text{O}_{(g)})$
 D. $\Delta H^\circ_r = [4 \text{ mol}](\Delta H^\circ_f \text{CO}_{2(g)}) + (5 \text{ mol})(\Delta H^\circ_f \text{H}_2\text{O}_{(g)}) - [2 \text{ mol}](\Delta H^\circ_f \text{C}_2\text{H}_{5(g)}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ_f \text{O}_{2(g)})$
 E. $\Delta H^\circ_r = [2 \text{ mol}](\Delta H^\circ_f \text{CO}_{2(g)}) + (\frac{4}{3} \text{ mol})(\Delta H^\circ_f \text{H}_2\text{O}_{(g)}) - [2 \text{ mol}](\Delta H^\circ_f \text{C}_2\text{H}_{5(g)}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ_f \text{O}_{2(g)})$
4. For which of the following statements is correct for the reaction:
 $2\text{H}_2\text{O}_{2(l)} \rightarrow 2\text{H}_2\text{O}_{(g)} + \text{O}_{2(g)} + 98.4 \text{ kJ}$
 A. exothermic, potential energy greater in reactants
 B. exothermic, potential energy greater in products
 C. endothermic, potential energy greater in products
 D. endothermic, potential energy less in reactants
 E. endothermic, potential energy less in products

Final Exam Review Chemistry 11 Answers

Mary Millhollon, Richard H. Langley



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Final Exam Review: Elementary Algebra A. A. Frempong, Elementary Algebra covers Signed Number and Real Number Operations Order of Operations and Evaluation of Expressions Exponential Notation and Rules of Exponents Polynomial addition subtraction multiplication and division Solving First Degree Equations Word Problems Ratio and Proportion Factoring Polynomials Solving quadratic equations by factoring Graphs Slopes Intercepts and Equations of Straight Lines Solving Systems of Linear Equations and Word Problems Radicals square roots addition Pythagorean Theorem and Applications Areas and Perimeters Algebraic Fractions reduction multiplication division Solving Linear inequalities

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choice items with constructed response formats This trend is expected to continue As a result a new item writing book is needed one that provides comprehensive coverage of both types of items and of the validity theory underlying them This book is an outgrowth of the author s previous book *Developing and Validating Multiple Choice Test Items* 3e Haladyna 2004 That book achieved distinction as the leading source of guidance on creating and validating selected response test items Like its predecessor the content of this new book is based on both an extensive review of the literature and on its author s long experience in the testing field It is very timely in this era of burgeoning testing programs especially when these items are delivered in a computer based environment Key features include Comprehensive and Flexible No other book so thoroughly covers the field of test item development and its various applications Focus on Validity Validity the most important consideration in testing is stressed throughout and is based on the Standards for Educational and Psychological Testing currently under revision by AERA APA and NCME Illustrative Examples The book presents various selected and constructed response formats and uses many examples to illustrate correct and incorrect ways of writing items Strategies for training item writers and developing large numbers of items using algorithms and other item generating methods are also presented Based on Theory and Research A comprehensive review and synthesis of existing research runs throughout the book and complements the expertise of its authors

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