

Instructions: On your scantron sheet enter and bubble in your name, UF ID number (start on the leftmost space and leave any extra spaces blank), and Form Code (see above). You may retain your exam sheet (mark your answers on it and the scantron sheet) and any other sheets. Turn in only your scantron.

Possibly useful info: $R=0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K}$; $R=8.314 \text{ J/mol}\cdot\text{K}$; $1 \text{ J} = 1 \text{ kg}\cdot\text{m}^2/\text{s}^2$; $h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$;
 $c = 2.998 \times 10^8 \text{ m/s}$; specific heat capacity (H_2O) = $4.184 \text{ J/g}\cdot^\circ\text{C}$

- Calculate the concentration (in molarity) of a NaOH solution if 25.0 mL of the solution are needed to neutralize 17.4 mL of a 0.312 M HCl solution.
 (1) 0.146 M (2) 0.217 M (3) 0.446 M (4) 0.517 M (5) 0.733 M
- A solid will float on any liquid that is more dense than the solid. The volume of a piece of calcite weighing 35.6 g is 12.9 cm^3 . On how many of the following liquids will calcite float? Carbon tetrachloride ($d=1.60 \text{ g/cm}^3$); methylene bromide ($d=2.50 \text{ g/cm}^3$); tetrabromoethane ($d=2.96 \text{ g/cm}^3$); methylene iodide ($d=3.33 \text{ g/cm}^3$)
 (1) 1 (2) 2 (3) 3 (4) 4 (5) 0
- What is the molecular formula of a gaseous compound (composed solely of C, H, and Cl atoms) whose mass is 0.568 g at 26°C and 740 torr in a 120 mL flask?
 (1) NH_4NO_3 (2) $\text{C}_2\text{H}_5\text{Cl}$ (3) $\text{C}_2\text{H}_4\text{Cl}_2$ (4) CHCl_3 (5) CCl_4
- For the XeF_4 molecule, what is the hybridization on Xe, the number of lone pairs (LP), and does it possess a dipole moment (DM) or not?
 (1) d^2sp^3 (or sp^3d^2), two LP, DM = 0 (2) sp^3 , zero LP, DM = 0
 (3) dsp^3 (or sp^3d), one LP, DM $\neq 0$ (4) dsp^3 (or sp^3d), two LP, DM $\neq 0$
 (5) d^2sp^3 (or sp^3d^2), two LP, DM $\neq 0$
- Which of the following would be predicted to have the smallest size and why?
 (1) Mg^{2+} because it has the most metallic character
 (2) Al^{3+} because it has the highest effective nuclear charge Z_{eff}
 (3) Ne because it is inert and has a filled shell
 (4) O^{2-} because it has the highest electron affinity
 (5) all of the above are the same size because they are all isoelectronic
- What is the root mean square speed of N_2 gas at STP?
 (1) 15.6 cm/s (2) 243 cm/s (3) $1.24 \times 10^4 \text{ cm/s}$ (4) $4.93 \times 10^2 \text{ cm/s}$ (5) 893 cm/s
- A liquid has a normal boiling point of 78°C and a vapor pressure of 400 Torr at 50°C . What is the heat of vaporization of this liquid?
 (1) 2.0 kJ/mol (2) 9.4 kJ/mol (3) 22 kJ/mol (4) 34 kJ/mol (5) 210 kJ/mol
- At 20°C , the vapor pressure of pure benzene is 75 torr and pure toluene is 22 torr. An equimolar solution of benzene and toluene is prepared. What is the mole fraction of benzene in the vapor above this solution at 20°C ?
 (1) 0.23 (2) 0.29 (3) 0.50 (4) 0.77 (5) 0.83
- A sample of liquid water is electrolyzed, completely converting the water to hydrogen gas and oxygen gas which is collected in a closed container. If the total pressure in the container is 600 torr, what is the partial pressure of the hydrogen in the container?
 (1) 100 torr (2) 200 torr (3) 300 torr (4) 400 torr (5) 500 torr
- The molecular shape of the IBr_3 molecule is
 (1) trigonal bipyramidal (2) seesaw (3) tee-shaped (4) trigonal pyramidal (5) trigonal planar
- Which of the following processes or reactions is not exothermic?
 (1) $\text{K}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{KBr}(\text{s})$ (2) $\text{CO}_2(\text{g}) \rightarrow \text{CO}_2(\text{s})$ (3) $\text{F}_2(\text{g}) \rightarrow 2\text{F}(\text{g})$
 (4) $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ (5) $\text{F}(\text{g}) + e^- \rightarrow \text{F}^-(\text{g})$
- Which of the following groups is arranged Left to Right in order of increasing atomic size?
 (1) Mg, Na, K (2) K, Ca, Sr (3) Na, Mg, Al (4) Rb, K, Na (5) Br, Cl, F
- A first-order reaction has a rate constant of $1.32 \times 10^{-3} \text{ yr}^{-1}$. What is the half-life of the reaction?
 (1) $1.32 \times 10^{-3} \text{ yr}$ (2) $1.90 \times 10^{-3} \text{ yr}$ (3) 525 yr (4) 758 yr (5) None of these

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Table of Contents Final Exam Chem 2045

1. Understanding the eBook Final Exam Chem 2045

- The Rise of Digital Reading Final Exam Chem 2045
- Advantages of eBooks Over Traditional Books
- 2. Identifying Final Exam Chem 2045
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Final Exam Chem 2045
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Final Exam Chem 2045
 - Personalized Recommendations
 - Final Exam Chem 2045 User Reviews and Ratings
 - Final Exam Chem 2045 and Bestseller Lists
- 5. Accessing Final Exam Chem 2045 Free and Paid eBooks
 - Final Exam Chem 2045 Public Domain eBooks
 - Final Exam Chem 2045 eBook Subscription Services
 - Final Exam Chem 2045 Budget-Friendly Options
- 6. Navigating Final Exam Chem 2045 eBook Formats
 - ePub, PDF, MOBI, and More
 - Final Exam Chem 2045 Compatibility with Devices
 - Final Exam Chem 2045 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Final Exam Chem 2045
 - Highlighting and Note-Taking Final Exam Chem 2045
 - Interactive Elements Final Exam Chem 2045
- 8. Staying Engaged with Final Exam Chem 2045
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Final Exam Chem 2045

9. Balancing eBooks and Physical Books Final Exam Chem 2045
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Final Exam Chem 2045
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Final Exam Chem 2045
 - Setting Reading Goals Final Exam Chem 2045
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Final Exam Chem 2045
 - Fact-Checking eBook Content of Final Exam Chem 2045
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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