



CHEMISTRY SSC-I

26

Time allowed: 2:40 Hours

Total Marks Sections B and C: 53

NOTE: Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separately provided answer book. Use supplementary answer sheet i.e. Sheet-B if required. Write your answers neatly and legibly.

SECTION – B (Marks 33)

Q. 2 Answer any ELEVEN parts. The answer to each part should not exceed 3 to 4 lines. (11 x 3 = 33)

- (i) Differentiate ion and free radical. Give one example of each.
- (ii) Ozone is a gas that stops UV radiations. Find the number of molecules in 3.5 moles of ozone gas.
- (iii) Differentiate shells and sub-shells.
- (iv) Define:
 - a. Periodic law
 - b. Group of Periodic Table
 - c. Atomic size
- (v) Briefly explain the formation of ions of Aluminium ${}_{13}^{27}\text{Al}$ and Fluorine ${}_{9}^{19}\text{F}$.
- (vi) How do the aqueous solutions of ionic compounds conduct electricity?
- (vii) Define Boyle's law. Derive its mathematical expression.
- (viii) What is meant by vapour pressure? Why does the vapour pressure of a liquid increase by increasing temperature?
- (ix) The molar mass of sodium hydroxide is 40 g/mol. If you dissolve 25 g of it to make 3 dm³ solution, what will be its molarity?
- (x) Two pairs of liquid are given. Which one is miscible and which one is not? State and give reason:
 - a. H_2O and CCl_4
 - b. H_2O and Methanol
- (xi) Separate oxidizing and reducing agents in the following reactions:
 - a. $\text{H}_2\text{S} + \text{Cl}_2 \rightarrow 2\text{HCl} + \text{S}$
 - b. $\text{Cu}^{+2}_{(\text{aq})} + \text{Zn}_{(\text{s})} \rightarrow \text{Zn}^{+2}_{(\text{aq})} + \text{Cu}_{(\text{s})}$
 - c. $2\text{Mg} + \text{CO}_2 \rightarrow 2\text{MgO} + \text{C}$
- (xii) Explain briefly the cathodic protection.
 - a. Why are alkali metals kept in kerosene oil?
 - b. Complete the following reactions:
 - (i). $2\text{Na} + 2\text{H}_2\text{O}_{(\text{l})} \rightarrow$
 - (ii). $\text{Mg} + \text{H}_2\text{O}_{(\text{g})} \rightarrow$
- (xiv) Write any three applications of Platinum.
- (xv) Write the electronic configuration of the following by using 1s, 2s..... formula:
 - a. ${}_{18}^{40}\text{Ar}$
 - b. ${}_{14}^{28}\text{Si}$
 - c. ${}_{7}^{14}\text{N}$

SECTION – C (Marks 20)

- Note:** Attempt any TWO questions. All questions carry equal marks. (2 x 10 = 20)
- Q. 3**
 - a. Write the postulates of Bohr atomic model. (05)
 - b. What is meant by shielding effect? Also write and give reason for its trends in periodic table. (1+4)
 - Q. 4**
 - a. What is ionic bond? Explain the formation of ionic bond in sodium chloride. (2+3)
 - b. Explain the effect of external pressure on boiling point. (05)
 - Q. 5**
 - a. Define solution, suspension and colloid. (03)
 - b. How is sodium manufactured from fused sodium chloride? Support your answer with equations and diagram. (3+2+2)

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