

(c) Standard enthalpy changes of reaction can also be determined using average bond enthalpies.

(i) What is meant by the term *average bond enthalpy*?

.....

.....

.....

..... [2]

Table 3.1 below shows some average bond enthalpies.

bond	average bond enthalpy / kJ mol^{-1}
C–H	+415
C–C	+345
C=C	+611

Table 3.1

(ii) Explain the bonding in a C=C double bond. Use the orbital overlap model.

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..... [2]

(iii) Suggest why the average bond enthalpy of a C=C bond is **not** twice the bond enthalpy of a C–C bond.

.....

.....

..... [1]

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 $360 \div 20 = 18$ angles of the quadrilateral are $3 \times 18, 5 \times 18, 9 \times 18, 13 \times 18$

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