

100) Carbonic acid reacts as a buffer at pH_a molecules. How H_2CO_3 molecules is a ring of eight carbon atoms many atoms of sulfur are there in a molecule of H_2CO_3 molecules?

$$\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 \rightleftharpoons \text{H}^+ + \text{HCO}_3^- \rightleftharpoons \text{H}^+ + \text{CO}_3^{2-}$$

$$\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 \rightleftharpoons \text{H}^+ + \text{HCO}_3^- \rightleftharpoons \text{H}^+ + \text{CO}_3^{2-}$$

101) The only intermediate forms in acid buffer are not the (H₂CO₃).

102) Describe how the acid buffer forms acid.

The acid buffer forms acid by the reaction of CO_2 and H_2O to form H_2CO_3 . The H_2CO_3 then reacts with H_2O to form H^+ and HCO_3^- . The H^+ then reacts with HCO_3^- to form H_2CO_3 .

103) Describe how the acid buffer forms acid buffer.

The acid buffer forms acid buffer by the reaction of CO_2 and H_2O to form H_2CO_3 . The H_2CO_3 then reacts with H_2O to form H^+ and HCO_3^- . The H^+ then reacts with HCO_3^- to form H_2CO_3 .

104) Carbonic acid is a component of acid buffer. Describe the reaction.

Carbonic acid reacts with water to form bicarbonate and hydrogen ions.

What is the reaction number of acid at $\text{H}_2\text{CO}_3/\text{HCO}_3^-/\text{H}^+$?

$$\text{H}_2\text{CO}_3 \rightleftharpoons \text{H}^+ + \text{HCO}_3^-$$



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