

- Instructions: (1) Substantiate your answers with formulas for finding using one or more of the following methods:
 (2) Balance each equation.
 (3) Identify the type reaction: synthesis, decomposition, single replacement, double replacement, or combustion.

1. Aluminum metal reacts with oxygen (air) to form aluminum oxide and oxygen diatomic molecules.

Type: synthesis

Reaction: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

2. Iron(III) sulfate solution and barium (OH) solution.

Type: double replacement

Reaction: $\text{Fe}_2(\text{SO}_4)_3 + 3\text{Ba}(\text{OH})_2 \rightarrow 3\text{BaSO}_4 + 2\text{Fe}(\text{OH})_3$

3. Calcium phosphate reacts with phosphoric acid to produce calcium dihydrogen phosphate. (Note: dihydrogen phosphate reactions that are not indicated are combustion.)

Type: double replacement

Reaction: $\text{Ca}_3(\text{PO}_4)_2 + 2\text{H}_3\text{PO}_4 \rightarrow 3\text{CaH}_2\text{P}_2\text{O}_7$

4. Dichlorophosphorus pentoxide is heated to form solid phosphorus pentoxide and oxygen gas.

Type: decomposition

Reaction: $\text{P}_2\text{O}_5 + 2\text{Cl}_2 \rightarrow 2\text{POCl}_3$

5. Dichlorophosphorus pentoxide reacts with sulfur trioxide (SO₃) to produce, in excess, the barium salt (see SO₃'s sulfate and sulfur oxides).

Type: double replacement

Reaction: $\text{P}_2\text{O}_5 + 3\text{SO}_3 \rightarrow 2\text{P}_2\text{S}_5$

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Highlights



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