

Background

- 1. Study of this \_\_\_\_\_ multi \_\_\_\_\_ relationship is a \_\_\_\_\_ method \_\_\_\_\_
- 2. Research on \_\_\_\_\_ followed \_\_\_\_\_ approaches \_\_\_\_\_

2.  $40g = 40r + 4 \cdot 40g$

2. Magnesium nitride reacts with a single oxygen atom to produce 2 formula units of magnesium oxide

The \_\_\_\_\_ is a \_\_\_\_\_ reaction \_\_\_\_\_ give the \_\_\_\_\_ \_\_\_\_\_ from the \_\_\_\_\_ involved in the reaction \_\_\_\_\_

Ex. Problem

Write balanced chemical equation with chemical states, chemical number is provided



moles reactants

$\frac{1}{2}$	$\text{Al}$	$\frac{3}{2}$	$\text{O}_2$
$\frac{1}{2}$	$\text{Al}$	$\frac{1}{2}$	$\text{Al}_2\text{O}_3$
$\frac{1}{2}$	$\text{O}_2$	$\frac{1}{2}$	$\text{Al}_2\text{O}_3$

2. If you start with 4 moles of Al, how many moles of  $\text{Al}_2\text{O}_3$  will be produced?

1 mol Al gives  $\text{Al}_2\text{O}_3$  1 mol  $\text{O}_2$  is 3 mol  $\text{Al}_2\text{O}_3$

# Episode 801 Chemistry Answer Key

**IM Harris**



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