

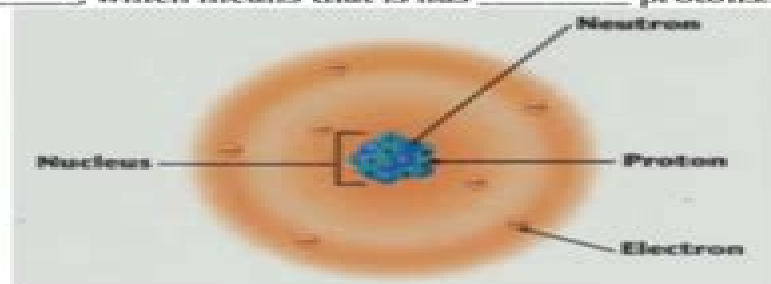
## Physical Science Standards:

**SPS1:** Students will investigate our current understanding of the atom.

### a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
  - Proton has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Protons are also the elements \_\_\_\_\_ number of the identity of the element, so protons are found by looking at the atomic number
  - Neutron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Neutrons are found by subtracting the \_\_\_\_\_ Number from the \_\_\_\_\_ Mass or the BIG #- SMALL #.
  - Electron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Electrons are equal to the number of \_\_\_\_\_ in a neutral atom.
  - The nucleus is made up of these 2 particles, \_\_\_\_\_ and \_\_\_\_\_.
  - The electrons are located in this region of the atom, known as the \_\_\_\_\_.
  - The electron cloud is divided in to energy levels or energy shells, the first level will hold \_\_\_\_\_ electrons and the second level will hold \_\_\_\_\_ electrons.
  - The valence shell or valence level is the \_\_\_\_\_ energy shell of the atom.
- atomic mass and atomic number.
  - Atomic number is equal to the number of \_\_\_\_\_ in the atom and is also the atoms identity—it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
  - Define Isotope \_\_\_\_\_.
  - Ex. Carbon-12 and Carbon-14
    - Carbon 12 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
    - Carbon 14 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
- explain the relationship of the proton number to the element's identity.
  - The \_\_\_\_\_ number is equal to the number of \_\_\_\_\_ in the atom which is how the element is identified.
  - Ex. Nitrogen has an atom number of \_\_\_\_\_, which means that is has \_\_\_\_\_ protons.

|            |                                              |
|------------|----------------------------------------------|
| molybdenum | ← element name                               |
| 42         | ← atomic number<br>number of protons (Z)     |
| Mo         | ← atomic symbol                              |
| 95.94      | ← atomic mass<br>A (this is an average mass) |



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

# Eoct Physics Study Guide

**Huangqi Zhang**



## **Eoct Physics Study Guide:**

## **Eoct Physics Study Guide** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the ability of words has are more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Eoct Physics Study Guide**, a literary masterpiece that delves deep in to the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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