

Chemistry EOC Review – Answer Key

- 1) a. 68,200,000 cg b. 5000 mm c. 0.0125 kL d. 0.548 dm
- 2) ~21.5 mL
- 3) production of heat & light, formation of a gas, formation of a precipitate, color change
- 4) $33.64 \text{ g} / 4.28 \text{ ml} = 7.86 \text{ g/ml} = \text{Fe}$
- 5) m.p. $-183^{\circ}\text{C} = \text{CH}_4$ (methane)
- 6) H_2 – splint test (burning splint ignites hydrogen)
 O_2 – splint test (oxygen reignites glowing splint)
 H_2O – cobalt chloride paper (paper changes from blue to pink in the presence of water vapor)
 CO_2 – splint test (carbon dioxide extinguishes burning splint);
 limewater test (limewater turns cloudy when carbon dioxide is bubbled through it,
 producing CaCO_3)
- 7) The protons are positively charged particles found inside the nucleus along with the neutrons, which have no charge. The total number of these nucleons in an atom is the mass # / atomic mass of the atom. The electrons are the negatively charged particles found orbiting in "clouds" around the nucleus. Electrons can be gained by the atom to form anions or lost to form cations.
- 8) The thing that makes any element on the periodic table different from all other elements is the number of protons in its nucleus. Isotopes are atoms of the same type of element which have different masses because they have different numbers of neutrons.
- 9) a. $^{39}_{19}\text{K}^{+}$
 b. $^{57}_{26}\text{Fe}$
 c. $^{33}_{16}\text{S}^{2-}$

<u>chemical symbol</u>	<u>atomic number</u>	<u>mass number</u>	<u># of protons</u>	<u># of electrons</u>	<u># of neutrons</u>	<u>net charge</u>
K ⁺	19	42	19	18	23	+1
Sr	38	88	38	38	50	0
F ⁻	9	20	9	10	11	-1
P	15	31	15	15	16	0
Ca ²⁺	20	42	20	18	22	+2

- 11) uranium - 238 $238 - 92 = 146$ neutrons
- 12) (1) 199.59g
(2) They contain the same # of protons (they are the same element), but different #s of neutrons (electrons don't matter, b/c they have virtually no mass), & thus different masses

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L Towner



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Table of Contents Eoc Review Quiz Answer Key Chemistry

1. Understanding the eBook Eoc Review Quiz Answer Key Chemistry
 - The Rise of Digital Reading Eoc Review Quiz Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Eoc Review Quiz Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Eoc Review Quiz Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Eoc Review Quiz Answer Key Chemistry
 - Personalized Recommendations
 - Eoc Review Quiz Answer Key Chemistry User Reviews and Ratings
 - Eoc Review Quiz Answer Key Chemistry and Bestseller Lists
5. Accessing Eoc Review Quiz Answer Key Chemistry Free and Paid eBooks
 - Eoc Review Quiz Answer Key Chemistry Public Domain eBooks
 - Eoc Review Quiz Answer Key Chemistry eBook Subscription Services
 - Eoc Review Quiz Answer Key Chemistry Budget-Friendly Options

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6. Navigating Eoc Review Quiz Answer Key Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Eoc Review Quiz Answer Key Chemistry Compatibility with Devices
 - Eoc Review Quiz Answer Key Chemistry Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Eoc Review Quiz Answer Key Chemistry
 - Highlighting and Note-Taking Eoc Review Quiz Answer Key Chemistry
 - Interactive Elements Eoc Review Quiz Answer Key Chemistry
 8. Staying Engaged with Eoc Review Quiz Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Eoc Review Quiz Answer Key Chemistry
 9. Balancing eBooks and Physical Books Eoc Review Quiz Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Eoc Review Quiz Answer Key Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Eoc Review Quiz Answer Key Chemistry
 - Setting Reading Goals Eoc Review Quiz Answer Key Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Eoc Review Quiz Answer Key Chemistry
 - Fact-Checking eBook Content of Eoc Review Quiz Answer Key Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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










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