

## Enzyme Action: Testing Catalase Activity

Many organisms can decompose hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) enzymatically. Enzymes are globular proteins, responsible for most of the chemical activities of living organisms. They act as *catalysts*, substances that speed up chemical reactions without being destroyed or altered during the process. Enzymes are extremely efficient and may be used over and over again. One enzyme may catalyze thousands of reactions every second. Both the temperature and the pH at which enzymes function are extremely important. Most organisms have a preferred temperature range in which they survive, and their enzymes most likely function best within that temperature range. If the environment of the enzyme is too acidic, or too basic, the enzyme may irreversibly *denature*, or unravel, until it no longer has the shape necessary for proper functioning.

$\text{H}_2\text{O}_2$  is toxic to most living organisms. Many organisms are capable of enzymatically destroying the  $\text{H}_2\text{O}_2$  before it can do much damage.  $\text{H}_2\text{O}_2$  can be converted to oxygen and water, as follows:



Although this reaction occurs spontaneously, enzymes increase the rate considerably. At least two different enzymes are known to catalyze this reaction: *catalase*, found in animals and protists, and *peroxidase*, found in plants. A great deal can be learned about enzymes by studying the rates of enzyme-catalyzed reactions. The rate of a chemical reaction may be studied in a number of ways including:

- measuring the rate of appearance of a product (in this case,  $\text{O}_2$ , which is given off as a gas)
- measuring the rate of disappearance of substrate (in this case,  $\text{H}_2\text{O}_2$ )
- measuring the pressure of the product as it appears (in this case,  $\text{O}_2$ ).



In this experiment, you will measure the rate of enzyme activity under various conditions, such as different enzyme concentrations, pH values, and temperatures. It is possible to measure the concentration of oxygen gas formed as  $\text{H}_2\text{O}_2$  is destroyed using an  $\text{O}_2$  Gas Sensor. If a plot is made, it may appear similar to the graph shown.

At the start of the reaction, there is no product, and the concentration is the same as the atmosphere. After a short time, oxygen accumulates at a rather constant rate. The slope of the curve at this initial time is constant and is called the *initial rate*. As the peroxide is destroyed, less of it is available to react and the  $\text{O}_2$  is produced at lower rates. When no more peroxide is left,  $\text{O}_2$  is no longer produced.

### OBJECTIVES

In this experiment, you will

- use an Oxygen Gas Sensor to measure the production of oxygen gas as hydrogen peroxide is destroyed by the enzyme catalase or peroxidase at various enzyme concentrations.
- measure and compare the initial rates of reaction for this enzyme when different concentrations of enzyme react with  $\text{H}_2\text{O}_2$ .

# Enzyme Action Testing Catalase Activity Lab Answers

**Mark Steyvers**



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