7.5 × 10⁻⁴. What is the value of equilibrium constant [RPMT-2001] (2) 1/3(4) None (1)3(3)1Q.21If the equilibrium constant for the reaction [RPMT-2002] N₂ + 3H₂ → 2NH₃, is K then the equilibrium, constant for the reaction $2N_{\star} + 6H_{\star} \rightarrow 4NH_{\star}$, would be equal to $(4) \frac{1}{122}$ (3) (2) JK (1) \mathbb{K}^{n} The reaction quotient (Q) for the reaction [CBSE-PMT-2003] $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ is given by $Q = \frac{[NH_3]^2}{[N_2][H_2]^3}$. The reaction will proceed from right to left if (1) $Q < K_s$ (2) $Q > K_o$ (3) Q = 0[CBSE-PMT-2003] The following equilibria are given Q.23 $N_2 + 3H_2 \rightarrow 2NH_3 \dots K_1$ $N_3 + O_2 \rightarrow 2NO \dots K_2$ $H_1 + \frac{1}{2}O_2 \rightarrow H_2O$ K_1 The equilibrium constant of the reaction $2NH_3 + \frac{5}{2}O_2 \rightarrow 2NO + 3H_2O$, in terms of $K_1 \cdot K_2$ and K_3 15 (1) K₁ K₂ / K₃ (2) K₁ K₂ / K₃ (3) K₂ K₃ / K₁ (4) K₁ K₂ K₃ For the gas phase reaction, 0.24 $C_1H_4 + H_2 \longrightarrow C_2H_4$ ($\Delta H = -32.7$ kcal). carried out in a vessel, the equilibrium conc. of C,H, can be increased by [AMU 2003] Decreasing the pressure (2) Increasing the temperature (3) Removing some C,H (4) Adding some H, Consider the following reversible reaction at equilibrium, $2H_1O(g) \Longrightarrow 2H_1(g) + O_1(g)$ Q.25 $\Delta H = 24.7$ kJ. Which one of the following changes in conditions will lead to maximum decomposition of H₂O(g)? [Kerala CEE 2004] (1) Increasing both temperature and pressure (2) Decreasing temperature and increasing pressure (3) Increasing temperature and decreasing pressure (4) Increasing temperature at constant pressure In the reaction, $A_{s}(g) + 4B_{s}(g) \rightleftharpoons 2AB_{s}(g) \Delta H < 0$, the formation of $AB_{s}(g)$ will be favoured at Q.26[MP PMT 2004] (1) High temperature, low pressure (2) Low temperature, high pressure (3) High temperature, high pressure (4) Low temperature, low pressure

For a reversible reaction forward and backward rate constant respectively are 2.5 × 10⁻⁴ and

Q.20

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